

### Lesson 3: Limiting reactants

- Reactions stop when one of the reactants is used up.

Limiting reactant - is used completely in the reaction; determines the amount of product formed

Excess reactant - leftover once the reaction stops.

- To calculate which reactant is limiting, you will compare the amount of each reactant to a single product (the true amount of product formed came from the limiting)



If 200.0 g of sulfur reacts with 100.0 g of chlorine, what mass of disulfur dichloride is produced? Which is the limiting reactant?

given = 200.0 g S<sub>8</sub>; 100.0 g Cl<sub>2</sub>  
unknown = ? g S<sub>2</sub>Cl<sub>2</sub>; ? limiting reactant

$$200.0 \text{ g S}_8 \times \frac{1 \text{ mol S}_8}{256.528 \text{ g S}_8} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol S}_8} \times \frac{135.038 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \\ = 421 \text{ g S}_2\text{Cl}_2$$

$$100.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{4 \text{ mol Cl}_2} \times \frac{135.038 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \\ = \boxed{190 \text{ g S}_2\text{Cl}_2}$$

Chlorine is the limiting reactant



If 100.0 g Na and 100.0 g Fe<sub>2</sub>O<sub>3</sub> react, what mass of Iron is produced? Which is the limiting reactant?

given = 100.0 g Na; 100.0 g Fe<sub>2</sub>O<sub>3</sub>  
unknown = ? g Fe; ? limiting reactant

$$100.0 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.990 \text{ g Na}} \times \frac{2 \text{ mol Fe}}{6 \text{ mol Na}} \times \frac{55.845 \text{ g Fe}}{1 \text{ mol Fe}} = \\ = 81.0 \text{ g Fe}$$

$$100.0 \text{ g Fe}_2\text{O}_3 \times \frac{1 \text{ mol Fe}_2\text{O}_3}{159.687 \text{ g Fe}_2\text{O}_3} \times \frac{2 \text{ mol Fe}}{1 \text{ mol Fe}_2\text{O}_3} \times \frac{55.845 \text{ g Fe}}{1 \text{ mol Fe}} = \\ = \boxed{70.0 \text{ g Fe}}$$

L.R. = Fe<sub>2</sub>O<sub>3</sub>

# Limiting Reactants Study Guide

**Iris Stovall**

## **Limiting Reactants Study Guide:**

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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## **Table of Contents Limiting Reactants Study Guide**

1. Understanding the eBook Limiting Reactants Study Guide
  - The Rise of Digital Reading Limiting Reactants Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reactants Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Study Guide
  - Personalized Recommendations

- Limiting Reactants Study Guide User Reviews and Ratings
- Limiting Reactants Study Guide and Bestseller Lists

5. Accessing Limiting Reactants Study Guide Free and Paid eBooks

- Limiting Reactants Study Guide Public Domain eBooks
- Limiting Reactants Study Guide eBook Subscription Services
- Limiting Reactants Study Guide Budget-Friendly Options

6. Navigating Limiting Reactants Study Guide eBook Formats

- ePUB, PDF, MOBI, and More
- Limiting Reactants Study Guide Compatibility with Devices
- Limiting Reactants Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Limiting Reactants Study Guide
- Highlighting and Note-Taking Limiting Reactants Study Guide
- Interactive Elements Limiting Reactants Study Guide

8. Staying Engaged with Limiting Reactants Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Limiting Reactants Study Guide

9. Balancing eBooks and Physical Books Limiting Reactants Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Limiting Reactants Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Limiting Reactants Study Guide

- Setting Reading Goals Limiting Reactants Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Limiting Reactants Study Guide

- Fact-Checking eBook Content of Limiting Reactants Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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