

Lesson 3: Limiting reactants

- Reactions stop when one of the reactants is used up.

Limiting reactant - is used completely in the reaction; determines the amount of product formed

Excess reactant - leftover once the reaction stops.

- To calculate which reactant is limiting, you will compare the amount of each reactant to a single product (the true amount of product formed came from the limiting)



If 200.0g of sulfur reacts with 100.0g of chlorine, what mass of disulfur dichloride is produced? which is the limiting reactant?

given = 200.0g S₈ ; 100.0g Cl₂

unknown = ?g S₂Cl₂ ; ? limiting reactant

$$200.0\text{g S}_8 \times \frac{1 \text{ mol S}_8}{256.528\text{g S}_8} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol S}_8} \times \frac{135.038\text{g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = 421\text{g S}_2\text{Cl}_2$$

$$100.0\text{g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906\text{g Cl}_2} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol Cl}_2} \times \frac{135.038\text{g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \boxed{190\text{g S}_2\text{Cl}_2}$$

Chlorine is the limiting reactant



If 100.0g Na and 100.0g Fe₂O₃ react, what mass of Iron is produced? which is the limiting reactant?

given = 100.0g Na ; 100.0g Fe₂O₃

unknown = ?g Fe ; ? limiting reactant

$$100.0\text{g Na} \times \frac{1 \text{ mol Na}}{22.990\text{g Na}} \times \frac{2 \text{ mol Fe}}{6 \text{ mol Na}} \times \frac{55.845\text{g Fe}}{1 \text{ mol Fe}} = 81.0\text{g Fe}$$

$$100.0\text{g Fe}_2\text{O}_3 \times \frac{1 \text{ mol Fe}_2\text{O}_3}{159.687\text{g Fe}_2\text{O}_3} \times \frac{2 \text{ mol Fe}}{1 \text{ mol Fe}_2\text{O}_3} \times \frac{55.845\text{g Fe}}{1 \text{ mol Fe}} = \boxed{70.0\text{g Fe}}$$

L.R. = Fe₂O₃

Limiting Reactants Study Guide

Iris Stovall



Limiting Reactants Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon: buckyballs, graphite, and diamond are illustrated at the left, as is the molecule methane CH_4 from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. **ASAP Chemistry: A Quick-Review Study Guide for the AP Exam** The Princeton Review, 2019-02-12 Looking for sample exams, practice questions, and test-taking strategies? Check out our extended in-depth AP chem prep guide: *Cracking the AP Chemistry Exam*. LIKE CLASS NOTES ONLY BETTER The Princeton Review's ASAP Chemistry is designed to help you zero in on just the information you need to know to successfully grapple with the AP test. No questions, no drills, just review. Advanced Placement exams require students to have a firm grasp of content; you can't bluff or even logic your way to a 5. Like a set of class notes borrowed from the smartest student in your grade, this book gives you exactly that. No tricks or crazy stratagems, no sample essays or practice sets. Just the facts, presented with lots of helpful visuals. Inside ASAP Chemistry, you'll find Essential concepts, terms, and functions for AP Chem, all explained clearly and concisely. Diagrams, charts, and graphs for quick visual reference. A three-pass icon system designed to help you prioritize learning what you MUST, SHOULD, and COULD know in the time you have available. Ask Yourself questions to help identify areas where you might need extra attention. A resource that's perfect for last-minute exam prep and for daily class work. Topics covered in ASAP Chemistry include Atomic structure, Covalent bonding, intermolecular

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