

## Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of  $\text{C}_2\text{H}_6$  and 3.44 g of  $\text{O}_2$ , determine the limiting reagent

$$14.8 \text{ g C}_2\text{H}_6 \times \frac{1 \text{ mol C}_2\text{H}_6}{26.06 \text{ g C}_2\text{H}_6} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_2\text{H}_6} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

$\text{O}_2$  is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.34 \text{ g H}_2\text{O}$$

c) determine the number of grams of  $\text{H}_2\text{O}$  produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol C}_2\text{H}_6}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g C}_2\text{H}_6}{1 \text{ mol C}_2\text{H}_6} = .0948 \text{ g C}_2\text{H}_6 \text{ used}$$

$$14.8 \text{ g C}_2\text{H}_6 - 0.0948 \text{ g C}_2\text{H}_6 = 13.9 \text{ g C}_2\text{H}_6 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{394.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol NaOH}}{241.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

# Limiting Reagent Worksheet 202 Answer Key

**G Thomas**

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## **Table of Contents Limiting Reagent Worksheet 202 Answer Key**

1. Understanding the eBook Limiting Reagent Worksheet 202 Answer Key
  - The Rise of Digital Reading Limiting Reagent Worksheet 202 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Worksheet 202 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent Worksheet 202 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Worksheet 202 Answer Key
  - Personalized Recommendations
  - Limiting Reagent Worksheet 202 Answer Key User Reviews and Ratings
  - Limiting Reagent Worksheet 202 Answer Key and Bestseller Lists
5. Accessing Limiting Reagent Worksheet 202 Answer Key Free and Paid eBooks
  - Limiting Reagent Worksheet 202 Answer Key Public Domain eBooks
  - Limiting Reagent Worksheet 202 Answer Key eBook Subscription Services
  - Limiting Reagent Worksheet 202 Answer Key Budget-Friendly Options

6. Navigating Limiting Reagent Worksheet 202 Answer Key eBook Formats
  - ePUB, PDF, MOBI, and More
  - Limiting Reagent Worksheet 202 Answer Key Compatibility with Devices
  - Limiting Reagent Worksheet 202 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reagent Worksheet 202 Answer Key
  - Highlighting and Note-Taking Limiting Reagent Worksheet 202 Answer Key
  - Interactive Elements Limiting Reagent Worksheet 202 Answer Key
8. Staying Engaged with Limiting Reagent Worksheet 202 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagent Worksheet 202 Answer Key
9. Balancing eBooks and Physical Books Limiting Reagent Worksheet 202 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent Worksheet 202 Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent Worksheet 202 Answer Key
  - Setting Reading Goals Limiting Reagent Worksheet 202 Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent Worksheet 202 Answer Key
  - Fact-Checking eBook Content of Limiting Reagent Worksheet 202 Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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