

# Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

## Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "n-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of substances.

## Model 1 – Ion Charges for Selected Elements

H <sup>+</sup>	Transition elements						Al <sup>3+</sup>	N <sup>3-</sup>	O <sup>2-</sup>	F <sup>1-</sup>
Li <sup>+</sup>	Be <sup>2+</sup>									
Na <sup>+</sup>	Mg <sup>2+</sup>	Fe <sup>2+</sup> Fe <sup>3+</sup>	Ni <sup>2+</sup> Ni <sup>3+</sup>	Cu <sup>+</sup> Cu <sup>2+</sup>	Zn <sup>2+</sup>			PO <sup>3-</sup>	S <sup>2-</sup>	Cl <sup>1-</sup>
K <sup>+</sup>	Ca <sup>2+</sup>									Br <sup>1-</sup>
Rb <sup>+</sup>	Sr <sup>2+</sup>			Ag <sup>1+</sup>			Sn <sup>2+</sup> Sn <sup>4+</sup>			I <sup>-</sup>
	Ba <sup>2+</sup>				Hg <sup>2+</sup> Hg <sup>4+</sup>		Pb <sup>2+</sup> Pb <sup>4+</sup>			

Cations

Anions

1. Based on the information in Model 1:

a. Identify three elements that form only one cation.

Cu<sup>+</sup> K<sup>+</sup> Na<sup>+</sup>

b. Identify three elements that form only one anion.

+ charged elements

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**Byte**, 1989-10

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