

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g } CS_2}{76.2\text{g } CS_2} \times \frac{1\text{mol } CS_2}{1\text{mol } CS_2} \times \frac{2\text{mol } SO_2}{3\text{mol } O_2} \times \frac{64.1\text{g } SO_2}{1\text{mol } SO_2} = 161\text{g } SO_2$$

$$\frac{100.\text{g } O_2}{32.0\text{g } O_2} \times \frac{1\text{mol } O_2}{3\text{mol } O_2} \times \frac{2\text{mol } SO_2}{1\text{mol } SO_2} \times \frac{64.1\text{g } SO_2}{1\text{mol } SO_2} = 134\text{g } SO_2$$

134 g of SO_2 can be produced in this reaction.

Limiting Reactant Quick Lab Answers

Neil D. Jespersen, Alison Hyslop

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