

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{ g CS}_2}{76.2\text{ g CS}_2} \times \frac{1\text{ mol CS}_2}{1\text{ mol CS}_2} \times \frac{2\text{ mol SO}_2}{1\text{ mol CS}_2} \times \frac{64.1\text{ g SO}_2}{1\text{ mol SO}_2} = 161\text{ g SO}_2$$

$$\frac{100.\text{ g O}_2}{32.0\text{ g O}_2} \times \frac{1\text{ mol O}_2}{3\text{ mol O}_2} \times \frac{2\text{ mol SO}_2}{1\text{ mol O}_2} \times \frac{64.1\text{ g SO}_2}{1\text{ mol SO}_2} = 134\text{ g SO}_2$$

134 g of SO₂ can be produced in this reaction.

Limiting Reactant Quick Lab Answers

Neil D. Jespersen, Alison Hyslop



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