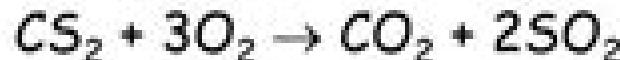


Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g } CS_2}{76.2\text{g } CS_2} \times \frac{1\text{mol } CS_2}{1\text{mol } CS_2} \times \frac{2\text{mol } SO_2}{3\text{mol } O_2} \times \frac{64.1\text{g } SO_2}{1\text{mol } SO_2} = 161\text{g } SO_2$$

$$\frac{100.\text{g } O_2}{32.0\text{g } O_2} \times \frac{1\text{mol } O_2}{3\text{mol } O_2} \times \frac{2\text{mol } SO_2}{1\text{mol } SO_2} \times \frac{64.1\text{g } SO_2}{1\text{mol } SO_2} = 134\text{g } SO_2$$

134 g of SO_2 can be produced in this reaction.

Limiting Reactants In A Recipe Answers

Tom Hsu PH.D.

Limiting Reactants In A Recipe Answers:

CCEA GCSE Double Award Science Denmour Boyd,Nora Henry,Frank McCauley,Alyn G. McFarland,James Napier,Roy White,2017-12-18 Exam Board CCEA Level GCSE Subject Science First Teaching September 2017 First Exam June 2019

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Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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