

Limiting Reactant

For the following reactions, find the following:

- Which of the reactants is the limiting reactant?
 - What is the maximum amount of each product that can be formed?
 - How much of the excess reactant is left over after the reaction is complete?
- See next page for calculations!

1. Consider the following reaction:



a) limiting reactant

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate.

$$\begin{aligned} & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 18.6 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{50.0 \text{ g Na}_3\text{PO}_4}{1} \times \frac{1 \text{ mol Na}_3\text{PO}_4}{162.94 \text{ g Na}_3\text{PO}_4} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol Na}_3\text{PO}_4} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 45.5 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{85.00 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 31.9 \text{ g NaNO}_3 \end{aligned}$$

2. Consider the following reaction:



a) limiting reactant

Answer the questions at the top of this sheet, assuming we start with 100.0 grams of calcium carbonate and 45.0 grams of iron (II) phosphate.

$$\begin{aligned} & \frac{100.0 \text{ g CaCO}_3}{1} \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCO}_3} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 103.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol FePO}_4} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 46.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Fe}_2(\text{CO}_3)_3}{2 \text{ mol FePO}_4} \times \frac{291.73 \text{ g Fe}_2(\text{CO}_3)_3}{1 \text{ mol Fe}_2(\text{CO}_3)_3} = 43.5 \text{ g Fe}_2(\text{CO}_3)_3 \end{aligned}$$

Limiting Reactant Packet

George W. Roberts



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Chemistry Rhobert Lewis, Wynne Evans, 2018-03-31 The fifth edition of this engaging and established textbook provides students with a complete course in chemical literacy and assumes minimal prior experience of science and maths Written in an accessible and succinct style this book offers comprehensive coverage of all the core topics in organic inorganic and physical chemistry Topics covered include bonding moles solutions and solubility energy changes equilibrium organic compounds and spectroscopy Each unit contains in text exercises and revision questions to consolidate learning at every step and is richly illustrated with diagrams and images to aid understanding This popular text is an essential resource for students who are looking for an accessible introductory textbook It is also ideal for non specialists on courses such as general science engineering environmental health or life sciences New to this Edition A foreword by Professor Sir John Meurig Thomas FRS former Director of the Royal Institution Three additional units on Gibbs Energy Changes Organic Mechanisms and Fire and Flame Accompanying online resources for this title can be found at bloomsburyonlineresources.com/chemistry/5e These resources are designed to support teaching and learning when using this textbook and are available at no extra cost

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Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science **Expert Systems in Chemistry Research** Markus C. Hemmer, 2007-12-13 Expert systems allow scientists to access manage and apply data and specialized knowledge from various disciplines to their own research Expert Systems in Chemistry Research explains the general scientific basis and computational principles behind expert systems and demonstrates how they can improve the efficiency of scientific workflows *Chemical Reactions and Chemical Reactors* George W. Roberts, 2008-03-14 Focused on the undergraduate audience Chemical Reaction Engineering provides students with complete coverage of the fundamentals including in depth coverage of chemical kinetics By introducing heterogeneous catalysis early in the book the text gives students the knowledge they need to solve real chemistry and industrial problems An emphasis on problem solving and numerical techniques ensures students learn and practice the skills they will need later on whether for industry or graduate work **The Encyclopedia of Physics** Robert Besancon, 2013-11-11 **General Chemistry** Kenneth W. Whitten, 1996 As in previous editions clearly explained problem solving strategies continue to be the strength of this student friendly text which can be viewed as a traditional general chemistry text Examples of this are that the text covers thermodynamics normally a topic split into two parts and covered in two different semesters in one chapter The book also follows a standard narrative example problem format has a solid traditional writing style and promotes problem solving However the authors have added elements to reflect changes in chemical education **Physics Briefs** , 1993 *Theoretical Chemistry in Belgium* Benoît Champagne, Michael S. Deleuze, Frank De Proft, Tom Leyssens, 2013-10-16 Readers of this volume can take a

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