

Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

Given the following reaction: $\text{C}_5\text{H}_{12} + 8 \text{O}_2 \rightarrow 5 \text{CO}_2 + 6 \text{H}_2\text{O}$

1. If 5.25 grams of pentane (C_5H_{12}) are completely reacted with oxygen (O_2), how many grams of carbon dioxide (CO_2) will be produced?

$$5.25 \text{ g C}_5\text{H}_{12} \times \frac{1 \text{ mol C}_5\text{H}_{12}}{72.151 \text{ g C}_5\text{H}_{12}} = 0.07276 \text{ mol C}_5\text{H}_{12}$$

$$0.07276 \text{ mol C}_5\text{H}_{12} \times \frac{5 \text{ mol CO}_2}{1 \text{ mol C}_5\text{H}_{12}} = 0.3638 \text{ mol CO}_2$$

$$0.3638 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 16.0 \text{ g CO}_2$$

2. If 9.75 grams of oxygen (O_2) are completely reacted with pentane (C_5H_{12}), how many grams of carbon dioxide (CO_2) will be produced?

$$9.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 0.3047 \text{ mol O}_2$$

$$0.3047 \text{ mol O}_2 \times \frac{5 \text{ mol CO}_2}{8 \text{ mol O}_2} = 0.1904 \text{ mol CO}_2$$

$$0.1904 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 8.38 \text{ g CO}_2$$

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of CO_2 actually form?

Limiting reactant: O_2

8.38 g CO_2 actually form

Limiting Reactants Answer Key

Ralph H. Petrucci, William S. Harwood



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Table of Contents Limiting Reactants Answer Key

1. Understanding the eBook Limiting Reactants Answer Key
 - The Rise of Digital Reading Limiting Reactants Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Answer Key
 - Personalized Recommendations
 - Limiting Reactants Answer Key User Reviews and Ratings
 - Limiting Reactants Answer Key and Bestseller Lists

5. Accessing Limiting Reactants Answer Key Free and Paid eBooks
 - Limiting Reactants Answer Key Public Domain eBooks
 - Limiting Reactants Answer Key eBook Subscription Services
 - Limiting Reactants Answer Key Budget-Friendly Options
6. Navigating Limiting Reactants Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactants Answer Key Compatibility with Devices
 - Limiting Reactants Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Answer Key
 - Highlighting and Note-Taking Limiting Reactants Answer Key
 - Interactive Elements Limiting Reactants Answer Key
8. Staying Engaged with Limiting Reactants Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Answer Key
9. Balancing eBooks and Physical Books Limiting Reactants Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Answer Key
 - Setting Reading Goals Limiting Reactants Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Answer Key
 - Fact-Checking eBook Content of Limiting Reactants Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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