

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{1 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 5.00 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = \boxed{19.87 \text{ g}}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answer Key

Paul B. Kelter



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Chemistry, Student Study Guide** James E. Brady, Fred

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Kelter, Steven S. Zumdahl, 2004-04 **Study Guide General Chemistry, Second Edition, Becker/Wentworth**
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Table of Contents Limiting Reactants Study Guide Answer Key

1. Understanding the eBook Limiting Reactants Study Guide Answer Key
 - The Rise of Digital Reading Limiting Reactants Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Study Guide Answer Key
 - Personalized Recommendations
 - Limiting Reactants Study Guide Answer Key User Reviews and Ratings
 - Limiting Reactants Study Guide Answer Key and Bestseller Lists
5. Accessing Limiting Reactants Study Guide Answer Key Free and Paid eBooks

- Limiting Reactants Study Guide Answer Key Public Domain eBooks
 - Limiting Reactants Study Guide Answer Key eBook Subscription Services
 - Limiting Reactants Study Guide Answer Key Budget-Friendly Options
6. Navigating Limiting Reactants Study Guide Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Limiting Reactants Study Guide Answer Key Compatibility with Devices
 - Limiting Reactants Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Limiting Reactants Study Guide Answer Key
 - Highlighting and Note-Taking Limiting Reactants Study Guide Answer Key
 - Interactive Elements Limiting Reactants Study Guide Answer Key
8. Staying Engaged with Limiting Reactants Study Guide Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Study Guide Answer Key
9. Balancing eBooks and Physical Books Limiting Reactants Study Guide Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Study Guide Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Study Guide Answer Key
- Setting Reading Goals Limiting Reactants Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Study Guide Answer Key
- Fact-Checking eBook Content of Limiting Reactants Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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