

Balancing Chemical Equations 10.3 - Answer Keys

Objectives: Use stoichiometry to identify the limiting and excess reactants.

Competencies: Assess the following balancing reaction problems. Show your work!

3. A researcher determines that 20.00 g of potassium iodide react with 10.00 g of silver nitrate to produce 18.00 g of silver iodide and an excess amount of potassium nitrate. How much potassium iodide was consumed by the following reaction and chemical equation?



a. Without the other reactants present, how much potassium iodide was consumed? How much potassium iodide reacted in the reaction?

$$10.00 \text{ g } AgNO_3 \times \left(\frac{1 \text{ mole } AgNO_3}{170.0 \text{ g } AgNO_3} \right) \left(\frac{1 \text{ mole } AgI}{170.0 \text{ g } AgI} \right) = \boxed{0.0588 \text{ mol}}$$

$$0.0588 \text{ mol } AgNO_3 \times \left(\frac{170.0 \text{ g } AgNO_3}{1 \text{ mole } AgNO_3} \right) \left(\frac{170.0 \text{ g } AgI}{170.0 \text{ g } AgI} \right) = \boxed{10.00 \text{ g } AgI}$$

b. Without the other reactants present?

Consumed (x)

a. Without the other reactants present, how much?

Consumed (y)

3. In a reaction, 15.00 g of magnesium sulfide reacts with an excess 20.00 g of aluminum sulfide.



a. Without aluminum sulfide, how much is consumed?

$$15.00 \text{ g } MgS \times \left(\frac{1 \text{ mole } MgS}{24.30 \text{ g } MgS} \right) \left(\frac{3 \text{ mole } Al_2S_3}{1 \text{ mole } MgS} \right) = \boxed{2.21 \text{ mol}}$$

$$2.21 \text{ mol } MgS \times \left(\frac{15.00 \text{ g } MgS}{1 \text{ mole } MgS} \right) \left(\frac{3 \text{ mole } Al_2S_3}{1 \text{ mole } MgS} \right) = \boxed{33.15 \text{ g}}$$

b. Without the other reactants present?

Consumed (x) + excess (10.00 g)

a. Without the other reactants present?

Excess Potassium Chloride (KClO3)

Limiting Reactants Study Guide Answer Key

Paul B. Kelter

Limiting Reactants Study Guide Answer Key:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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