

JEE Advanced 2020 Chemistry | Paper-2 | Code-E

SECTION 1 (Maximum Marks: 18)

- This section contains **SIX (06)** questions.
- The answer to each question is a **SINGLE DIGIT INTEGER** ranging from **0 TO 9, BOTH INCLUSIVE**.
- For Each question, enter the correct integer corresponding to the answer using the mouse and the on-screen virtual numeric keypad in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme:
Full Marks : +3 If ONLY the correct numerical value is entered;
Zero Marks : 0 If the question is unanswered;
Negative Marks : -1 In all other cases.

1. The 1st, 2nd, and the 3rd ionization enthalpies, I_1 , I_2 , and I_3 , of four atoms with atomic numbers n , $n + 1$, $n + 2$, and $n + 3$, where $n < 10$, are tabulated below. What is the value of n ?

Atomic number	Ionization Enthalpy (kJ/mol)		
	I_1	I_2	I_3
n	1681	3374	6050
$n + 1$	2081	3952	6122
$n + 2$	496	4562	6910
$n + 3$	738	1451	7733

Answer: 9

Solution:

According to the tabulated data,

Element with Atomic number ($n + 2$), should be an alkali metal

As we see, first ionization enthalpy (I_1) is very less but second ionization enthalpy (I_2) is very large.

Hence, atomic number can be = 11

That is
$$(n + 2) = 11$$

$$n = 9$$

Note: 'n' can't be '1'.

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