



Java Apple Manual

Ben Greisler

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that run on the iPhone iPod touch and iPad You'll learn how to take advantage of the iPad's larger display and the new features of iOS 4.2 By the end of this book you'll be ready to create the next hit in the iTunes App Store Designed for Easy Learning Key Skills Concepts Chapter opening lists of specific skills covered in the chapter Try This Hands on exercises that show you how to apply your skills Notes Extra information related to the topic being covered Tips Helpful reminders or alternate ways of doing things Cautions Errors and pitfalls to avoid Annotated Syntax Example code with commentary that describes the programming techniques being illustrated **iPhone SDK Programming: A Beginner's Guide** James

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was produced in what mass of iodine was produced

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times 2 \text{ mole } H_2O / 2 \text{ mole } H_2S = 16.9 \text{ g } H_2O$ is your theoretical yield so percent yield $137.1 \text{ g } H_2O / 149.4 \text{ g } H_2O \times 100 = 91.77\%$ problems 6-8 are similar to problem 5

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web 7 $\text{Na}_2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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