

16. (c) 350 cm, convergent
(d) 150 cm, divergent

17. Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C:

(a) 2 amp (b) 1 amp
(c) 3 amp (d) 0.6 amp

18. The device used to generate electrical energy is:
(a) Electric motor (b) Generator
(c) Galvanometer (d) Voltmeter

19. The number of molecules in 11 g of CO_2 are:
(a) 0.25×10^{23} (b) 0.50×10^{23}
(c) 1.00×10^{23} (d) 1.51×10^{23}

20. The value of charge / mass ratio of electron was determined by:
(a) W.K. Roentgen (b) J.J. Thomson
(c) Marie Curie (d) Niels Bohr

21. Which of the following reactions is an example of combination reaction?
(a) $2\text{Zn}(\text{s}) + \text{CuSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu}(\text{s})$
(b) $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaCl}(\text{aq})$
(c) $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$
(d) $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$

22. The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula:
(a) XY (b) XY_3
(c) XY_2 (d) X_2Y_3

23. The atomic radius (pm) of Li, Na, K and Rb varies in the order:
(a) $\text{Na} < \text{K} < \text{Rb} < \text{Li}$ (b) $\text{K} < \text{Na} < \text{Li} < \text{Rb}$
(c) $\text{Li} < \text{Na} < \text{K} < \text{Rb}$ (d) $\text{Rb} < \text{K} < \text{Na} < \text{Li}$

24. The electron affinity value (kJ mol^{-1}) of fluorine (F) is less than:
(a) Hydrogen (H) (b) Lithium (Li)
(c) Oxygen (O) (d) Chlorine (Cl)

25. Which of the following is an example of strong electrolyte?
(a) H_2CO_3 (b) NH_4OH
(c) NaCl (d) $\text{COOH} - \text{COOH}$

26. If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be:
(a) two times (b) five times
(c) ten times (d) twenty times

27. Bleaching powder is manufactured by the reaction of:
(a) CaCl_2 and CaCO_3 (b) Cl_2 and $\text{Ca}(\text{OH})_2$

28. (c) Cl_2 and CaSO_4 (d) Cl_2 and $\text{Ca}(\text{HCO}_3)_2$
The hard glass is obtained by fusing:
(a) soda ash, sand and limestone
(b) a mixture of sand, lime, borax and alkali carbonates
(c) potassium carbonate and limestone
(d) potassium carbonate, lead oxide and sand

29. Which of the following metals is most reactive?
(a) Aluminium (b) lead
(c) Mercury (d) Silver

30. Which of the following alloys contains chromium?
(a) Steel (b) Stainless steel
(c) Magnalium (d) Brass

31. Which of the following is monomer of natural rubber?
(a) Chloroethene (b) Chloroprene
(c) Isoprene (d) Buta-1, 3-diene

32. Alkaline KMnO_4 oxidises propanone to:
(a) propanoic acid (b) ethanoic acid
(c) methanolic acid (d) oxalic acid

33. Which of the following metals can displace Zn from ZnSO_4 solution?
(a) Calcium (b) Copper
(c) Iron (d) Mercury

34. Heating of sodium ethanoate with soda lime yields:
(a) ethane (b) methane
(c) ethanol (d) methanol

35. The compound formed by the reaction of ethyne with bromine is:
(a) $\text{Br}-\text{CH}=\text{CH}-\text{Br}$ (b) $\text{Br}-\text{CH}_2-\text{CHBr}_2$
(c) $\text{CH}_2=\text{CH}-\text{Br}$ (d) $\text{Br}_2\text{CH}-\text{CHBr}_2$

36. Cell were first discovered by Robert Hooke in the year:
(a) 1665 (b) 1674 (c) 1831 (d) 1839

37. The plant cells have a rigid cell wall that lies:
(a) Outside the plasma membrane
(b) Inside the plasma membrane
(c) In between the plasma membranes
(d) None of the above

38. During mitosis nucleolus and nuclear membrane are lost in stage:
(a) Prophase (b) Metaphase
(c) Anaphase (d) Telophase

39. An example of simple and permanent tissue is:
(a) Xylem (b) Phloem
(c) Sclerenchyma (d) All the above

40. The shape of squamous epithelial tissue is:
(a) Cubical (b) Flattened
(c) Pillar like (d) None of the above

Fresh water sponge belongs to phylum:
(a) Arthropoda (b) Annelida

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