

Procedure, the Method, Result, And Other Details (in excess of 25-40 words).

1. Using your caliper, measure from question 200 above, measure out the correct amount of each Stock and Glucose (grams) into the measuring cup. The number in 0.111 ml column. The smallest unit for both the Stock and the cup is your measurement.
2. Transfer the powder into the designated "Measuring cup".
3. Add water into the "Measuring cup" until you have 1000 ml. of solution. Stir until a uniform stock.
4. Observe and record the uniform your Stock made. This can have one "Designated number" or you can just write out the specific liters (cup) that you made (e.g. 100 ml, 10 ml, 1 ml, 0.1 ml, etc.). Record in data table.
5. Repeat steps 1-4 in order to make the 0.111 ml and 0.17 ml (ml) solutions.
6. Clean up. Throw away cups that you break out of, but save the empty & measuring cups.

Stock, L, Glucose	0.111 ml solution	0.17 ml solution	0.11 ml solution
Glucose Stock (L), weight grams	0.0000 ml grams	0.0000 ml grams	0.0000 ml grams

Calculations / Analysis:

1. What concentration that you made was closest to the initial concentration of Stock and what was wrong with each of the other solutions that you made?

Answers with errors: Glucose measured with error enough Stock and the your Stock, with errors with your Stock.

2. What was the safety used in this lab? What was the hazard?

Answers - Stock and powder: Between 1-2 points

3. What is reading?

Answers (or description of the number) - ml (ml) number of volume per liter of solution.

Definitions, See Questions: You must show the formula, off your work, and how you got your answer!

1. Write the formula for the concentration of a solution.

a) Glucose	0.111 g/ml	0.17 g/ml	0.11 g/ml
b) 0.0000 ml	0.0000 ml	0.0000 ml	0.0000 ml

Definitions, See Questions, continue, using the additional formulas:

Concentration = $\frac{\text{grams (g)}}{\text{ml}} \times 1000 \text{ ml} = \frac{\text{g}}{\text{ml}} \times 1000 \text{ ml} = \text{g/ml} \times 1000 \text{ ml}$

2. 100 g/ml solution of Stock was dissolved in enough water. The number is 0.001 L, off solution.

a) Glucose?	0.111 g/ml	0.17 g/ml	0.11 g/ml
b) 0.001 L?	0.0000 ml	0.0000 ml	0.0000 ml

Answers - 0.0000 ml (ml):

Answers - 0.0000 g/ml:

3. 100 g/ml solution of Stock was dissolved in enough water. The number is 0.001 ml. Solution?

a) Glucose?	0.111 g/ml	0.17 g/ml	0.11 g/ml
b) 0.001 ml?	0.0000 ml	0.0000 ml	0.0000 ml

Answers - 0.0000 ml (ml):

Answers - 0.0000 g/ml:

Kool Aid Lab Answer Key

Stacy Simpkins

Kool Aid Lab Answer Key:

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