

Limiting Reactants Practice



a) What is the limiting reactant if 0.866 mol of NO is mixed with 0.503 mol of O₂ to create NO₂?

$$0.866 \text{ mol NO} \times \frac{2 \text{ mol NO}_2}{2 \text{ mol NO}} = 0.866 \text{ mol NO}_2$$

limiting reactant product made

$$0.503 \text{ mol O}_2 \times \frac{2 \text{ mol NO}_2}{1 \text{ mol O}_2} = 1.006 \text{ mol NO}_2$$

b) Calculate the number of moles of NO₂ produced. (Answer should already be solved from the previous questions)

0.866 mol NO₂



a) If 7.40 g of O₃ reacts with 0.670 g of NO, which compound will be the limiting reactant?

O₃
O: 3 × 15.999 = 47.997 g/mol
⇒ 1 mol O₃ = 47.997 g O₃

$$7.40 \text{ g O}_3 \times \frac{1 \text{ mol O}_3}{47.997 \text{ g O}_3} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol O}_3} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 7.093 \text{ g NO}_2$$

NO
N: 1 × 14.007 = 14.007
O: 1 × 15.999 = 15.999
⇒ 1 mol NO = 30.006 g/mol

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol NO}} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 1.027 \text{ g NO}_2$$

limiting reactant product made

b) How many grams of NO₂ will be produced? (Answer should already be solved from the previous questions)

1.027 g NO₂

NO₂
N: 1 × 14.007 = 14.007
O: 2 × 15.999 = 31.998
⇒ 1 mol NO₂ = 46.005 g/mol

c) Calculate mass, in grams, of the excess reactant remaining at the end of the reaction. How many moles of excess reactant would this be?

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol O}_3}{1 \text{ mol NO}} \times \frac{47.997 \text{ g O}_3}{1 \text{ mol O}_3} = 1.072 \text{ g O}_3 \text{ used!}$$

$$\underbrace{7.40 \text{ g O}_3}_{\text{original}} - \underbrace{1.072 \text{ g O}_3}_{\text{used}} = \underline{6.328 \text{ g O}_3 \text{ leftover}}$$

Limiting Reactants Key

John A. Olmsted, Gregory M. Williams



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