



What are the properties of light?

Background

Light is both a wave and a particle. The wave properties of light include frequency and wavelength. On an atomic level light also shows particle-like properties. A photon is a 'particle' of light. Within the visible spectrum, we see different wavelengths of light as different colors. The primary colors are red, green, and blue. Other colors can be made by mixing red, green, and blue. Blue light is more energetic than red light. The investigation will use photo luminescence to demonstrate the different energy content of red and blue light. A diffraction grating splits light up into its spectrum.

Almost all the light you see is produced by atoms. When atoms absorb energy, electrons rise to higher energy levels. When the electrons fall back to their lower energy state, they may release energy in the form of light. In some elements it takes time for the energized electrons to fall back and give up their energy. These elements store energy and give off light slowly over a period of time. This is how glow-in-the-dark material works. Embedded in the material are atoms of the element phosphorus. When light energy hits the phosphorus atoms, some of the electrons absorb energy. When the electrons fall back, they release the stored energy and the material glows. The glow stops when all the electrons have returned to the lowest energy level. The process is called **photoluminescence**. The word "photo" means light and the word "luminescence" means glowing.

The quantum theory of light

- A photon is a small quantity (like a particle) of light. You can think of a photon like a short packet of a wave. The intensity of light describes how many photons per second are produced (or absorbed). The color of each photon depends on its energy. Different colors of light are produced by photons with different energies. High energy photons have shorter wavelengths than low energy photons.
- In order to glow, an electron in a phosphorus atom must first absorb energy from a photon of light. One of the predictions of the quantum theory is that electrons cannot have any energy, but can have only the specific energies corresponding to the energy levels in a particular atom. That means a photon must have enough energy to boost the electron up one whole level to be absorbed. If the photon does not have enough energy, the electron cannot absorb it and the phosphorus atom cannot glow.

Virtual vs Real Images Can you tell what type of image a lens will make? Virtual and real images are two kinds of images that are apparent reproduction of real objects formed by a mirror or a lens. These images are formed through processes of reflection, refraction or diffracted rays of light. In a real image, light rays are brought to a focus at the position of the image. The most important aspect of image is that a real image can be made visible on a screen such as a sheet of paper whereas a virtual image cannot be produced on a screen. Virtual images are made by rays that do not actually come from where one sees the image.

In brief:

- While real image appears inverted, virtual image appears erect
- While real image can be obtained on a screen, virtual image cannot be seen on a screen.
- In case of mirrors, real image lies in front while virtual image lies behind the mirror.
- In case of lenses, real image lies on the other side of the object whereas virtual image lies on the same side of the object.

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Leading Questions

- What creates light? _____
- How does glow in the dark material work? _____
- What creates the different colors of light? _____
- How do we make different colors of light? _____
- How can we tell what colors of light we are seeing? _____

Light Lab Answer Sheet

Karen C. Timberlake



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