

**What are the properties of light?****Background**

Light is both a wave and a particle. The wave properties of light include frequency and wavelength. On an atomic level light also shows particle-like properties. A photon is a 'particle' of light. Within the visible spectrum, we see different wavelengths of light as different colors. The primary colors are red, green, and blue. Other colors can be made by mixing red, green, and blue. Blue light is more energetic than red light. The investigation will use photo luminescence to demonstrate the different energy content of red and blue light. A diffraction grating splits light up into its spectrum.

Almost all the light you see is produced by atoms. When atoms absorb energy, electrons rise to higher energy levels. When the electrons fall back to their lower energy state, they may release energy in the form of light. In some elements it takes time for the energized electrons to fall back and give up their energy. These elements store energy and give off light slowly over a period of time. This is how glow-in-the-dark material works. Embedded in the material are atoms of the element phosphorus. When light energy hits the phosphorus atoms, some of the electrons absorb energy. When the electrons fall back, they release the stored energy and the material glows. The glow stops when all the electrons have returned to the lowest energy level. The process is called photoluminescence. The word "photo" means light and the word "luminescence" means glowing.

**The quantum theory of light**

- A photon is a small quantity (like a particle) of light. You can think of a photon like a short packet of a wave. The intensity of light describes how many photons per second are produced (or absorbed). The color of each photon depends on its energy. Different colors of light are produced by photons with different energies. High energy photons have shorter wavelengths than low energy photons.
- In order to glow, an electron in a phosphorus atom must first absorb energy from a photon of light. One of the predictions of the quantum theory is that electrons cannot have any energy, but can have only the specific energies corresponding to the energy levels in a particular atom. That means a photon must have enough energy to boost the electron up one whole level to be absorbed. If the photon does not have enough energy, the electron cannot absorb it and the phosphorus atom cannot glow.

**Virtual vs. Real Images** Can you tell what type of image a lens will make? Virtual and real images are two kinds of images that are apparent reproduction of real objects formed by a mirror or a lens. These images are formed through processes of reflection, refraction or diffracted rays of light. In a real image, light rays are brought to a focus at the position of the image. The most important aspect of image is that a real image can be made visible on a screen such as a sheet of paper whereas a virtual image cannot be produced on a screen. Virtual images are made by rays that do not actually come from where one sees the image.

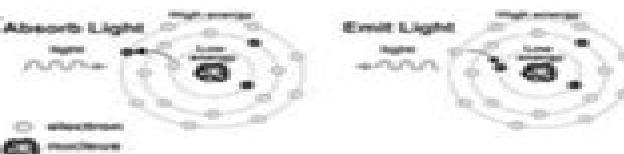
**In brief:**

- While real image appears inverted, virtual image appears erect.
- While real image can be obtained on a screen, virtual image cannot be seen on a screen.
- In case of mirrors, real image lies in front while virtual image lies behind the mirror.
- In case of lenses, real image lies on the other side of the object whereas virtual image lies on the same side of the object.

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**Leading Questions:**

- What creates light? \_\_\_\_\_
- How does glow in the dark material work? \_\_\_\_\_
- What creates the different colors of light? \_\_\_\_\_
- How do we make different colors of light? \_\_\_\_\_
- How can we tell what colors of light we are seeing? \_\_\_\_\_



# Light Lab Answer Sheet

**Karen C. Timberlake**

## **Light Lab Answer Sheet:**

**Exploring Physical Science in the Laboratory** John T. Salinas ,2019-02-01 This full color manual is designed to satisfy the content needs of either a one or two semester introduction to physical science course populated by nonmajors It provides students with the opportunity to explore and make sense of the world around them to develop their skills and knowledge and to learn to think like scientists The material is written in an accessible way providing clearly written procedures a wide variety of exercises from which instructors can choose and real world examples that keep the content engaging Exploring Physical Science in the Laboratory guides students through the mysteries of the observable world and helps them develop a clear understanding of challenging concepts

**Geology From Experience** E. Kirsten Peters,Larry E. Davis,2000-11-05 Moving away from the observation and vocabulary focus of traditional physical geology lab manuals Peters and Davis s Geology from Experience offers experiments that favor hands on involvement and scientific problem solving Students are asked to use geological tools and techniques analyze data from observation experiment and research solve simple equations and make assessments and relevant predictions This approach class tested with great success by the authors gives students a real taste of the scientific experience by revealing the ways geologists actually do their work

**Hands-On Physical Science** Laurie E. Westphal,2021-09-09 Hands On Physical Science immerses students in the world of real life chemists and physicists Through engaging authentic learning experiences students will engage in fascinating experiments while building STEM skills This book is packed with activities that can easily be conducted in the classroom using everyday materials and includes everything teachers need to help students think critically and problem solve as they explore the fascinating world of physical science From examining Newton s laws using sports video clips to studying energy through the design and building of roller coasters students will not just learn about physical science they will be scientists Grades 6 8

***Life in the Laboratory*** Donald Glen Humphrey,Henry Van Dyke,David L. Willis,1965

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**Phy:P&P** Zitzewitz,1998-07

***Life in the Lab*** Humphrey,1997-11

## **Laboratory Manual for Chemistry**

Karen Timberlake,1979

***Biological Investigations Lab Manual*** Warren Dolphin,David Vleck,Linda Westgate,James Colbert,2010-01-27 The lead author of eight successful previous editions has brought together a team that combined has well over 60 years experience in offering beginning biology labs to several thousand students each year at Iowa State University Their experience and diverse backgrounds ensure that this extensively revised edition will meet the needs of a new generation of students Designed to be used with all majors level general biology textbooks the included labs are investigative using both discovery and hypothesis based science methods Students experimentally investigate topics observe structure use critical thinking skills to predict and test ideas and engage in hands on learning Students are often asked what evidence do you have that in order to encourage them to think for themselves By emphasizing investigative quantitative and comparative approaches to the topics the authors continually emphasize how the biological sciences are integrative yet unique An

instructor's manual available through McGraw Hill Lab Central provides detailed advice based on the authors experience on how to prepare materials for each lab teachings tips and lesson plans and questions that can be used in quizzes and practical exams This manual is an excellent choice for colleges and universities that want their students to experience the breadth of modern biology **Laboratory Manual for Chemistry** Karen C. Timberlake,1976 **Laboratory Manual for Optical Mineralogy** Norris W. Jones,Fred Donald Bloss,1980

**Introductory Astronomy Laboratory Manual** Donald J. Summers,1997 **Laboratory Manual to Accompany Chemistry in Context** American Chemical Society,2005-02 The 5th edition Laboratory Manual that accompanies Chemistry in Context is compiled and edited by Gail Steehler Roanoke College The experiments use microscale equipment wellplates and Beral type pipets as well as common materials Project type and cooperative collaborative laboratory experiments are included Additional experiments are available on the Online Learning Center as is the instructor's guide **The Filmmaker's Handbook** Steven Ascher,1999 [Light and Lighting](#),1956

*Filming for Television* A. Arthur Englander,Paul Petzold,1976 **Modules** ,2005 **Introduction to Chemistry, Laboratory Manual** T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students **Instruction papers**

International correspondence schools,1909 **The Extraordinary Chemistry of Ordinary Things, Laboratory Manual** Bruce Coburn Richardson,Carl H. Snyder,Thomas G. Chasteen,1995-01-12 Using ordinary and several not so ordinary products as examples this book explores the chemical principles behind them to show how chemistry affects our daily lives It includes an environmental chapter that focuses on pollution and its effects It also examines how these chemical principles affect our lives on a larger scale **Pulp and Paper Magazine of Canada** ,1920

## **Light Lab Answer Sheet** Book Review: Unveiling the Power of Words

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