

Gas Law Study Guide

1. A gas is initially at 200 kPa pressure and 27°C. If the temperature is increased by 10 degrees Celsius, what will the new pressure be if volume remains constant?
2. The volume of a gas is 1.1 L and at 300 K. Assuming that the volume is 10 L, what is the temperature of the gas?
3. A known balanced reaction: $\text{Ca} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2$ (Note: mass of H_2 was for produced from completely reacting 2.0 g of Ca at 20°C and 1.0 atm pressure)
4. A 0.5 liter 2.0 L of gas at a pressure of 0.5 atm and a temperature of 100°C, what will be the new pressure of the gas if I decrease the volume of the gas to 2.0 L and the pressure is 1.0 atm?
5. Use the reaction above to calculate the mass of iron that must be used to obtain 0.100 L of hydrogen gas at STP. The chemical equation for the reaction is: $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$

Gas Law Study Guide

**Carl David Morneweck, Chester Arthur
Buckner, Gale R. Kirschner**

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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993 *Fundamentals of Chemistry, Study Guide* James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

Study Guide with ActivPhysics Alan Van Heuvelen, Richard Wolfson, Jay M. Pasachoff, 1999 *Chemistry, Study Guide* James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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Understanding Physics David C. Cassidy,Gerald Holton,F. James Rutherford,2013-11-27 Understanding Physics provides a thorough grounding in contemporary physics while placing physics into its social and historical context Based in large part on the highly respected Project Physics Course developed by two of the authors it also integrates the results of recent

pedagogical research The text thus teaches about the basic phenomena in the physical world and the concepts developed to explain them shows that science is a rational human endeavor with a long and continuing tradition involving many different cultures and people develops facility in critical thinking reasoned argumentation evaluation of evidence mathematical modeling and ethical values The treatment emphasizes not only what we know but also how we know it why we believe it and what effects that knowledge has Why do we believe the Earth and planets revolve around the Sun Why do we believe that matter is made of atoms How do relativity theory and quantum mechanics alter our conception of Nature and in what ways do they leave the classical concepts unchanged What impact does the knowledge of finite energy resources have on our society How have applications of fundamental science such as the steam engine the laser the electric generator the transistor affected our lives How does the evidence for non scientific ideas such as UFOs ESP and the like differ from the evidence for accepted scientific results *Oil and Gas Law in a Nutshell* John S. Lowe, 1983

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