

Specific Heat Capacity Practice Problems

Name:

1. When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$C_p, w = Q/m \cdot \Delta T$$
$$Q = C_p, w (m) \Delta T$$
$$= (4186)(3)(10 - 80)$$
$$= 879,060 \text{ J}$$

2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30.0°C to 150°C?

$$Q = C_p, a (m_a) (\Delta T_a)$$
$$= 899 (.3) (150 - 30)$$
$$= 432,364 \text{ J}$$

3. Calculate the temperature change when:

a) 10.0 kg of water loses 232 kJ of heat.

$$232 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 232,000 \text{ J}$$
$$-232,000 \text{ J} = (4186)(10.0 \text{ kg}) \Delta T$$
$$-232,000 \text{ J} = 41860 \Delta T$$

b) 1.96 kJ of heat are added to 500. g of copper.

$$1.96 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 1960 \text{ J}$$
$$500 \text{ g} \cdot \frac{1 \text{ kg}}{1000 \text{ g}} = .5 \text{ kg}$$
$$1960 \text{ J} = (387)(.5 \text{ kg}) \Delta T$$
$$\frac{1960 \text{ J}}{193.5 \text{ J}} = \frac{193.5 \text{ J}}{10.1 \text{ °C}} = \Delta T$$

4. $2.52 \times 10^4 \text{ J}$ of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

$$T_f = 130^\circ\text{C}$$

$$Q = 2.52 \times 10^4 \text{ J}$$

$$m = 2.0 \text{ kg}$$

$$C_p, m = 138$$

$$2.52 \times 10^4 = (138)(2)(130 - T_i)$$

$$2.52 \times 10^4 = 276(130 - T_i)$$

$$91.3 = 130 - T_i$$

$$-38.7 = -T_i$$
$$38.7^\circ\text{C} = T_i$$

Heat Practice Problems Answers

WJ Hussar

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