

CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

a. What is the limiting reagent, and what is the reactant in excess?

b. Calculate the mass of FeS formed.

2. Acrylonitrile, $\text{C}_3\text{H}_3\text{N}$, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C_3H_6 , by reaction with nitric oxide, NO , as follows:



What mass of $\text{C}_3\text{H}_3\text{N}$ can be made when 21.6 g of C_3H_6 react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction: $\text{P}_{(s)} + 6 \text{Cl}_{(g)} \rightarrow 4 \text{PCl}_3_{(s)}$
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid, HCHO_2 , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.

a. Balance the following reaction: $\text{Zn}_{(s)} + \text{HCl}_{(aq)} \rightarrow \text{ZnCl}_2_{(aq)} + \text{H}_2_{(g)}$

b. A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl_2 (grams)	H_2 (L)
Before reaction				
After reaction	1.26 g			

Limiting Reactant Problems Answer Key

John A. Olmsted, Gregory M. Williams

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon buckyballs, graphite, and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course. The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy. The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry. Topical problem solving will focus on limiting reactant stoichiometry, identifying types of chemical reactions, equilibrium, acid base equilibria, and electrochemistry. Why would this resource be of interest to chemistry students? To be successful to get into a well known college, medical school, physical therapy, or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course. To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail. This is the weeding out process that every pre health student is apprehensive about. To succeed in this competitive environment entails not just studying harder or longer, it means re organizing textbook content so that it is meaningful to the student. This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career.

Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills, calculation methods and the language and basic concepts of chemistry. The book is designed to meet the specific needs of underprepared students. Concepts are presented only as they are needed and developed from the simple to the complex. The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter, energy and measurement, the properties of atoms, description of chemical bonding, study of chemical change and nuclear and organic chemistry. Undergraduate students will find the book as a very valuable academic material.

Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition, *Investigating Chemistry: A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the

fundamental chemical concepts most often covered in an introductory nonmajors course. Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students. This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package.

Prentice Hall Chemistry, 2000 Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject. Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications. Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry, organic chemistry and biochemistry. An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career.

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Student Solutions Manual: Ssm Chemistry Deborah Wiegand, 2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems. Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems.

Holt Chemistry Ralph Thomas Myers, 2004

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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999 Illinois Chemistry Teacher, 1992 **The Science Teacher**, 1992 Some issues are accompanied by a CD ROM on a selected topic Chemical Principles William L. Masterton, Emil J. Slowinski, Conrad L. Stanitski, 1981 Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a

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