



Topic 19.1: Organic chemistry – Fundamentals of organic chemistry

Organic chemistry focuses on the chemistry of compounds containing carbon.

• **Understanding:** A homologous series is a series of compounds of the same family, with the same general formula, which differ from each other by a common structural unit.

- **Homologous series:** series of compound of the same family, with the same general formula, which differ by a common structural unit
 - Common structure unit: often CH_2 (methylene) group
 - Same family: has the same functional group
 - Properties: homologous series has similar physical and chemical properties

• **Understanding:** Structural formulas can be represented in full and condensed format.

	Full structural formula	Condensed structural formula	Skeletal structural formula
Example		$\text{CH}_3\text{—CH}_2\text{—CH}_2\text{—CH}_3$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	
Description	Two-dimensional representation showing all atoms, their relative positions and bonds	Two-dimensional representation showing all atoms and their relative position.	Two dimensional representation of a compound where each vertex and end of line represent a carbon atom, and functional group is added

• **Understanding:** Structural isomers are compounds with the same molecular formula but different arrangements of atoms.

- **Structural isomers:** compounds with same molecular formula with different arrangement of atoms
 - Properties: structural isomers have different chemical and physical properties (melting point, boiling point etc.)

• **Understanding:** Functional groups are the reactive parts of molecules.

- **Functional groups:** reactive parts of a molecule that is responsible for the molecule's chemical property
 - Naming: the position of the function group is identified by the number of carbon to which it is attached

• **Understanding:** Saturated compounds contain single bonds only and unsaturated compounds contain double or triple bonds.

- **Saturated compounds:** organic compounds in which all the carbon-carbon bonds are single (e.g. **alkanes**)
- **Unsaturated compounds:** organic compound in which there is at least one carbon-carbon bond that is double or triple (e.g. **alkenes**)

• **Understanding:** Benzene is an aromatic, unsaturated hydrocarbon.

- **Aliphatic:** characterized by a straight, branched or cyclic structure
- **Aromatic:** characterized by presence of benzene ring
- **Benzene:** aromatic, unsaturated hydrocarbon in which six carbons bonds into a ring
 - **Sigma (σ) bond in benzene:** benzene contains six sp^2 hybridized carbon atoms bond to one other by sigma bonds
 - **Pi (π) bond in benzene:** p-orbitals in carbon overlap to form a continuous π bond above and below the plane of carbons

• **Applications and skills:** Explanation of the trends in boiling points of members of a homologous series.

Name	Formula	Structural formula (condensed)	Boiling point ($^{\circ}\text{C}$)
Methane	CH_4	CH_4	-161
Ethane	C_2H_6	CH_3CH_3	-89
Propane	C_3H_8	$\text{CH}_3\text{CH}_2\text{CH}_3$	-42
Butane	C_4H_{10}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	-0.5
Pentane	C_5H_{12}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	36
Hexane	C_6H_{14}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	69

- **Trend in boiling points:** boiling points increases as the carbon chain increases by a common structural unit
 - Trends arise from increasing intermolecular (dispersion forces) with increased chain lengths

• **Applications and skills:** Distinction between empirical, molecular and structural formulas.

- **Empirical formula:** simplest ratio of atoms (of an element) in a molecule
- **Molecular formula:** actual number of atoms (of each element) in a molecule
- **Structural formula:** representation of the type of atoms and their position relative to one other

Ib Chemistry Hl Level Paper 23

Chao Zhang

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Journal of Research of the National Bureau of Standards United States. National Bureau of Standards, 1974

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Section 5. Chemistry in relation to agriculture and applied botany. Section 6. Chemistry in relation to applied zoology and veterinary science. Section 7. Chemistry in relation to food and nutrition, 1950

Proceedings of the XIth International Congress of Pure and Applied Chemistry, 1947

Nuclear Science Abstracts, 1961

Popular Photography, 1995-01

Cumulated Index Medicus, 1966

Bibliographies and Literature of Agriculture, 1978

Journal of the American Society of Mechanical Engineers American Society of Mechanical Engineers, 1918

Canadian

Chemical Processing ,1972 American Journal of Physics ,1984 Archival journal targeted toward advanced level physics and physics education with its focus on the teaching and cultural aspects of physics **Sea-level Fluctuation and Coastal Evolution** William Armstrong Price,1987 **Actes** ,1972 **Science** ,1959 Scientific American ,1897 Proceedings of the Sixth International Congress on Catalysis, Imperial College, London, 12-16 July, 1976 Geoffrey Colin Bond,Peter B. Wells,Frederick Clifford Tompkins,1977 **The Cumulative Book Index** ,1965 A world list of books in the English language *Pollution* Royal Society of Chemistry (Great Britain),1990 Much has happened in the pollution field since 1982 when the first edition of this work was written This edition includes new subject areas such as radioactive pollution and chemistry and pollution of the stratosphere

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