

5

ELECTRONS IN ATOMS**SECTION 5.1 MODELS OF THE ATOM (pages 127–132)**

This section summarizes the development of atomic theory. It also explores the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Development of Atomic Models (pages 127–128)

1. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

2. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false

► The Bohr Model (pages 128–129)

3. What is a small, discrete unit of energy called?

It is called a quantum.

4. Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to

- a. move an electron from its present energy level to the next lower one
- b. maintain an electron in its present energy level

c. move an electron from its present energy level to the next higher one

5. In general, the higher the electron is on the energy ladder, the

further it is from the nucleus.

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Life Study Guide David E. Sadava, Gordon H. Orians, Craig Heller, William K. Purves, 2006-12-22 Especially helpful for AP Biology students each chapter of the study guide offers a variety of study and review tools The contents of each chapter are broken down into both a detailed review of the Important Concepts covered and a boiled down Big Picture snapshot The guide also covers study strategies common problem areas and provides a set of study questions both multiple choice and short answer

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Table of Contents Introduction To Atoms Guided Reading And Study Answers

1. Understanding the eBook Introduction To Atoms Guided Reading And Study Answers
 - The Rise of Digital Reading Introduction To Atoms Guided Reading And Study Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Introduction To Atoms Guided Reading And Study Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Introduction To Atoms Guided Reading And Study Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Introduction To Atoms Guided Reading And Study Answers
 - Personalized Recommendations
 - Introduction To Atoms Guided Reading And Study Answers User Reviews and Ratings
 - Introduction To Atoms Guided Reading And Study Answers and Bestseller Lists
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 - Introduction To Atoms Guided Reading And Study Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Introduction To Atoms Guided Reading And Study Answers
 - Highlighting and Note-Taking Introduction To Atoms Guided Reading And Study Answers
 - Interactive Elements Introduction To Atoms Guided Reading And Study Answers
8. Staying Engaged with Introduction To Atoms Guided Reading And Study Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Introduction To Atoms Guided Reading And Study Answers

9. Balancing eBooks and Physical Books Introduction To Atoms Guided Reading And Study Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Introduction To Atoms Guided Reading And Study Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Introduction To Atoms Guided Reading And Study Answers

- Setting Reading Goals Introduction To Atoms Guided Reading And Study Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Introduction To Atoms Guided Reading And Study Answers

- Fact-Checking eBook Content of Introduction To Atoms Guided Reading And Study Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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