

### MOLE Practice Problems

1. Atomic Mass vs. Molar Mass: Calculate the molar mass/ atomic mass of the following atoms and compounds.

	List atom types & AT's of each type of atom	atomic mass (with correct units)	Molar mass (with correct units)
$\text{H}_2\text{O}$	H = 2 O = 1	18.02 amu	18.02 grams/ mole
$\text{Fe}_2\text{O}_3$	Fe = 2 O = 3	159.70 amu	159.70 g/ mol
NaCl	Na = 1 Cl = 1	58.44 amu	58.44 g/mol
$\text{CO}_2$	C = 1 O = 2	44.01 amu	44.01 g/ mol
$\text{Ba}(\text{OH})_2$	Ba = 1 O = 2 H = 2	171.35 amu	171.35 g/mol

#### Conceptual Understanding Questions

2. How is the male similar to a deer?

**It's a word that represents an amount or a number**

3. Which contains more atoms, a mole of copper or a mole of gold? (write down for in each one)

They both contain  $6.02 \times 10^{23}$  atoms

4. Which weighs more, a mole of copper or a mole of gold? (write down how much each one weighs)

Gold = 196.97 g/mol      Copper = 63.55 g/mol

Q. How are the terms "molar mass" and "atomic mass" different from each other?

**Molar Mass = Mass of 1 mol ( $6.02 \times 10^{23}$ ) of a substance (g/mol)**

Atomic Mass = Mass of 1 atom (amu)

b. How are the terms the same?

Both the same number on the periodic table, you just need to change the units (g/mol -vs- amu)

6. Write down the conversion factors that would be used to convert between the following units:

Grams ← ..... → MOLES ← ..... → ATOMS & MOLECULES

1 mole / M grams      1 mole =  $6.02 \times 10^{23}$  atoms & molecules

**Conversion Problems:** For each problem you must use Dimensional Analysis & box your answer.

7. Given 12 g of aluminum, how many moles do you have?

$$\frac{12 \text{ g-Al}}{26.98 \text{ g-Al}} \times 1 \text{ mol Al} = 0.44 \text{ mol Al}$$

16. Given 3.4 moles  $\text{CF}_3$ , how many grams would you have?

$$34 \text{ mol } \text{CF}_4 \quad 86 \text{ g } \text{CF}_4 \quad = 2992 \\ 1 \text{ mol } \text{CF}_4 \quad = 3.0 \times 10^3 \text{ g } \text{CF}_4$$

# Mixed Practice Mole Problems Answers Worksheet

## Answer

**John Olmsted, Greg Williams, Robert C. Burk**

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## **Table of Contents Mixed Practice Mole Problems Answers Worksheet Answer**

1. Understanding the eBook Mixed Practice Mole Problems Answers Worksheet Answer
  - The Rise of Digital Reading Mixed Practice Mole Problems Answers Worksheet Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Practice Mole Problems Answers Worksheet Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mixed Practice Mole Problems Answers Worksheet Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Practice Mole Problems Answers Worksheet Answer
  - Personalized Recommendations
  - Mixed Practice Mole Problems Answers Worksheet Answer User Reviews and Ratings
  - Mixed Practice Mole Problems Answers Worksheet Answer and Bestseller Lists
5. Accessing Mixed Practice Mole Problems Answers Worksheet Answer Free and Paid eBooks
  - Mixed Practice Mole Problems Answers Worksheet Answer Public Domain eBooks
  - Mixed Practice Mole Problems Answers Worksheet Answer eBook Subscription Services
  - Mixed Practice Mole Problems Answers Worksheet Answer Budget-Friendly Options

6. Navigating Mixed Practice Mole Problems Answers Worksheet Answer eBook Formats
  - ePUB, PDF, MOBI, and More
  - Mixed Practice Mole Problems Answers Worksheet Answer Compatibility with Devices
  - Mixed Practice Mole Problems Answers Worksheet Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mixed Practice Mole Problems Answers Worksheet Answer
  - Highlighting and Note-Taking Mixed Practice Mole Problems Answers Worksheet Answer
  - Interactive Elements Mixed Practice Mole Problems Answers Worksheet Answer
8. Staying Engaged with Mixed Practice Mole Problems Answers Worksheet Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mixed Practice Mole Problems Answers Worksheet Answer
9. Balancing eBooks and Physical Books Mixed Practice Mole Problems Answers Worksheet Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mixed Practice Mole Problems Answers Worksheet Answer
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mixed Practice Mole Problems Answers Worksheet Answer
  - Setting Reading Goals Mixed Practice Mole Problems Answers Worksheet Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mixed Practice Mole Problems Answers Worksheet Answer
  - Fact-Checking eBook Content of Mixed Practice Mole Problems Answers Worksheet Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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