

**MOLE Practice Problems****1. Atomic Mass vs. Molar Mass:** Calculate the molar mass/ atomic mass of the following atoms and compounds:

	List atom types & #'s of each type of atom	atomic mass (with correct units)	Molar mass (with correct units)
H <sub>2</sub> O	H = 2      O = 1	18.02 amu	18.02 grams/ mole
Fe <sub>2</sub> O <sub>3</sub>	Fe = 2    O = 3	159.70 amu	159.70 g/ mol
NaCl	Na = 1    Cl = 1	58.44 amu	58.44 g/mol
CO <sub>2</sub>	C = 1    O = 2	44.01 amu	44.01 g/ mol
Ba(OH) <sub>2</sub>	Ba = 1    O = 2    H = 2	171.35 amu	171.35 g/mol

**Conceptual Understanding Questions**

2. How is the mole similar to a dozen?

It's a word that represents an amount or a number

3. Which contains more atoms, a mole of copper or a mole of gold? (write down # in each one)

They both contain  $6.02 \times 10^{23}$  atoms

4. Which weighs more, a mole of copper or a mole of gold? (write down how much each one weighs)

Gold = 196.97 g/mol

Copper = 63.55 g/mol

5. a. How are the terms "molar mass" and "atomic mass" different from each other?

Molar Mass = Mass of 1 mol ( $6.02 \times 10^{23}$ ) of a substance (g/mol)Atomic Mass = Mass of 1 atom (amu)

b. How are the terms the same?

Both the same number on the periodic table, you just need to change the units (g/mol -vs- amu)

6. Write down the conversion factors that would be used to convert between the following units:

Molar Mass (periodic table)                      Avogadro's #  
**GRAMS**  $\xleftarrow{\text{1 mole / \# grams}}$  **MOLES**  $\xrightarrow{\text{1 mole = } 6.02 \times 10^{23} \text{ atoms \& molecules}}$  **ATOMS \& MOLECULES**

**Conversion Problems:** For each problem you must use Dimensional Analysis & box your answer.

7. Given 12 g of aluminum, how many moles do you have?

$$\frac{12 \text{ g-Al}}{26.98 \text{ g-Al}} \times \frac{1 \text{ mol Al}}{1 \text{ mol Al}} = 0.44 \text{ mol Al}$$

16. Given 34 moles CF<sub>4</sub>, how many grams would you have?

$$\frac{34 \text{ mol-CF}_4}{1 \text{ mol-CF}_4} \times \frac{88 \text{ g CF}_4}{1 \text{ mol-CF}_4} = 2992 \text{ g CF}_4 = 3.0 \times 10^3 \text{ g CF}_4$$

# Mixed Practice Mole Problems Answers Worksheet Answer

**John Olmsted, Greg Williams, Robert C.  
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
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