



- 1) What is happening to the average kinetic energy of the molecules in the sample during section 2? **remains constant**
- 2) As a substance goes through section (2), what happens to the distance between the particles? **increases**
- 3) What is the name of the process happening during section (4)? **boiling (vaporization)**
- 4) What would be the name of the process happening during section (4) if time were going the other way? **condensation**
- 5) What is the melting point of this substance? **0 °C**
- 6) At what temperature would this sample finish boiling? **100 °C**
- 7) When this substance is melting, the temperature of the ice-water mixture remains constant because:
 - d. Heat energy is being converted to kinetic energy**
- 8) When a given quantity of water is heated at a constant rate, the phase change from liquid to gas takes longer than the phase change from solid to liquid because
 - a. The heat of vaporization is greater than the heat of fusion**
- 9) The temperature at which a substance in the liquid state freezes is the same as the temperature at which the substance
 - a. Melts**
- 10) Is this curve showing an **endothermic** or an **exothermic** process?
- 11) 34°C is equal to **307 K**.
- 12) -128°C is equal to **145 K**.

Heating Curve Problems Key

L Cohen

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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999 [Catalog of National Bureau of Standards Publications, 1966-1976: Citations and abstracts](#) United States. National Bureau of Standards. Technical Information and Publications Division, 1978 **Publications of the National Bureau of Standards ... Catalog** United States. National Bureau of Standards, 1975 **Publications of the National Institute of Standards and Technology ... Catalog** National Institute of Standards and Technology (U.S.), 1976 **NBS Special Publication**, 1968 **Catalog of National Bureau of Standards Publications, 1966-1976** United States. National Bureau of Standards, 1978

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