

Problem Set: Calculating Heat

Although the calorimeters are used to measure heat, the fact is that heat cannot be measured directly. Instead, the calorimeter is used to measure the change in temperature of a measured amount of a given substance. Heat is then calculated from this data; **heat is a derived quantity**. Recall that the textbook defines specific heat capacity (c) as the quantity of heat required to raise the temperature of a gram of the substance 1°C . The equation is as follows.

$$\text{specific heat} = \frac{\text{quantity of heat}}{\text{mass} \times \text{temperature change}}$$

The equation used to calculate heat from experimental data can be obtained from the above equation simply by rearranging terms. Multiplying both sides of the equation by (mass \times temperature change) gives:

$$q = mc\Delta T$$

where q = quantity of heat involved; m = mass; c = specific heat of substance; ΔT = change in temp.

To solve problems involving heat calculations, simply rearrange terms to isolate the unknown variable.

Example A

How much heat will be absorbed by 320 g of water when its temperature is raised by 35°C ? The specific heat of water is $4.18 \text{ J/(g} \times ^{\circ}\text{C)}$.

Solution

$$q = mc\Delta T$$

$$q = (320\text{g})(4.18 \frac{\text{J}}{\text{g} \cdot ^{\circ}\text{C}})(35^{\circ}\text{C})$$

$$q = 46,816 \text{ J} = \underline{47,000\text{J or } 47 \text{ kJ}} \text{ (only 2 sig figs should be in answer)}$$

Example B

Calculate the specific heat for aluminum if 16,500 J of heat are absorbed in raising the temperature of $1.50 \times 10^2 \text{ g}$ of aluminum by 125°C .

Solution

$$q = mc\Delta T$$

$$16,500 \text{ J} = (1.5 \times 10^2 \text{ g})(c)(125^{\circ}\text{C})$$

$$c = 0.880 \frac{\text{J}}{\text{g} \cdot ^{\circ}\text{C}}$$

Now you try it!

1. How much heat will be given off by 55 g of water as it cools from 87°C to 25°C ?

$$\Delta T = T_f - T_i = 25^{\circ}\text{C} - 87^{\circ}\text{C} = -62^{\circ}\text{C}$$

$$q = mc\Delta T$$

$$q = (55\text{g})(4.18 \frac{\text{J}}{\text{g} \cdot ^{\circ}\text{C}})(-62^{\circ}\text{C})$$

$$q = -14,000\text{J} = \underline{-14\text{kJ}} \text{ (exothermic)}$$

2. Calculate the specific heat of glass from the following data. The temperature of a piece of glass with a mass of 65 grams increases by 26°C when it absorbs 840 J of heat energy.

$$q = mc\Delta T$$

$$840\text{J} = (65\text{g})(c)(26^{\circ}\text{C})$$

$$c = \underline{0.50 \frac{\text{J}}{\text{g} \cdot ^{\circ}\text{C}}}$$

Heat Calculations Answers

William Perren Maycock

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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