

# Heat Problems I

## Standards:

- 3.1.10 B Describe the concepts of models as a way to predict and understand science and technology.
- 3.4.10 B Analyze energy sources and transfers of heat.

## Specific Heat Capacities of some Common Substances

Substance	$C_p$ (J/g°C)	Substance	$C_p$ (J/g°C)
Water (l)	4.184	Iron (s)	0.45
Water (s)	2.03	Mercury (l)	0.14
Water (g)	2.0	Carbon (s)	0.71
Aluminum (s)	0.89	Silver (s)	0.24
		Gold (s)	0.13

$$1\text{Cal} = 4.184\text{Joules}$$

$$1\text{kJ} = 1000\text{J}$$

## Problems:

- Convert the following number of calories into joules. (Use D.A. and sig figs.)
  - 100.0 cal = \_\_\_\_\_ J

- $1.00 \times 10^3$  cal = \_\_\_\_\_ J

- Convert the following numbers of joules (J) into kilojoules (kJ). (Use D.A. and sig figs.)

- 243,000 J = \_\_\_\_\_ kJ

- 0.251 J = \_\_\_\_\_ kJ

# Heat And Temperature Answer Sheet

**Ying Liu**

## **Heat And Temperature Answer Sheet:**

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