

## Section A

Answer all questions in the spaces provided

0 | 1

Anhydrous magnesium chloride,  $MgCl_2$ , can absorb water to form the hydrated salt  $MgCl_2 \cdot 4H_2O$



*Creating crystals directly is not really possible.*

0 | 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

*Virtually impossible to prevent dissolving.*

0 | 1 - 2

Some enthalpies of solution are shown in Table 1.

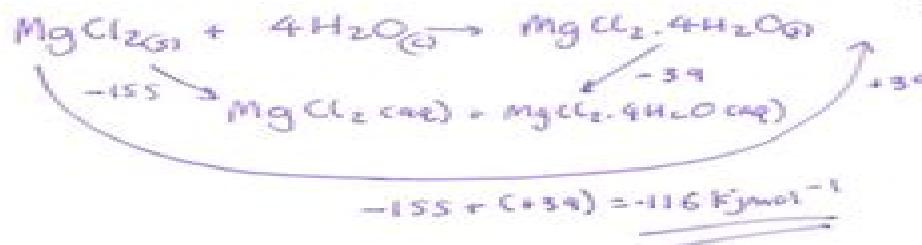
Table 1

*making the salts aqueous!*

Salt	Enthalpy of solution / $\text{kJ mol}^{-1}$
$MgCl_2(s)$	-155
$MgCl_2 \cdot 4H_2O(s)$	-39

Calculate the enthalpy change for the absorption of water by  $MgCl_2(s)$  to form  $MgCl_2 \cdot 4H_2O(s)$ .

[2 marks]



Enthalpy change \_\_\_\_\_  $\text{kJ mol}^{-1}$



0 2

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