

20. Which of the following correctly represents the ionization of an atom?

- a) $\text{Cl(g)} + e^- \rightarrow \text{Cl}^-(g)$
- b) $\text{Na(g)} \rightarrow \text{Na}^+(g) + e^-$
- c) $\text{Na(s)} - e^- \rightarrow \text{Na}^+(g)$
- d) $\text{Cl}_2(g) \rightarrow 2 \text{Cl(g)}$

21. Which of the following is likely to have the largest atomic radius?

~~at H~~ ~~at Mn~~ ~~at Cl~~ ~~at Rb~~ ~~at Ag~~
most low \rightarrow LEFT

22. Which one of the following isoelectronic species has the smallest radius? ~~most protons~~

- a) Mg^{2+} $12p^0$
- b) Na^+ $11p^0$
- c) Ne $10p^0$
- d) F^- $9p^0$
- e) O^{2-} $8p^0$

NOTE: explain your reasoning on the last page.

23. Which of the following has the greatest ionization energy? ~~smallest~~

- a) K
- b) Ca
- c) Fe
- d) Ga
- e) Br

~~all period 4~~

24. Which of the following has the lowest ionization energy? ~~largest~~

- a) Li
- b) Na
- c) K
- d) Rb
- e) Cs

~~same family~~

25. The successive ionization energies for one of the period three elements is listed below. Which element is referred to?

E_1	577.4 kJ/mol
E_2	1,816 kJ/mol
E_3	2,744 kJ/mol
E_4	11,580 kJ/mol
E_5	15,030 kJ/mol

valency electrons

- a) Na
- ~~b) Mg~~
- c) Al
- d) Si
- e) P

NOTE: explain your reasoning on the last page.

Multiple Choice Answers:

Use CAPITAL LETTERS for your answers...

- 1. B
- 2. D
- 3. C
- 4. D
- 5. C
- 6. E
- 7. A
- 8. C
- 9. B
- 10. B
- 11. D
- 12. B
- 13. C
- 14. D
- 15. D
- 16. E
- 17. A
- 18. A
- 19. A
- 20. B
- 21. D
- 22. A
- 23. E
- 24. E
- 25. ~~C~~

26. Draw the orbital diagram for a neutral Ag atom:
~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~
~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~
~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~

26. ~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~
27. ~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~
28. ~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~

29. ~~5s 5p 3d 4f~~ ~~5s 5p 3d 4f 47e-~~

Write the electron configuration for silver:

27. Long form:
 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2$

28. Short form:
 $[Ar] 4d^{10} 5s^1$

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