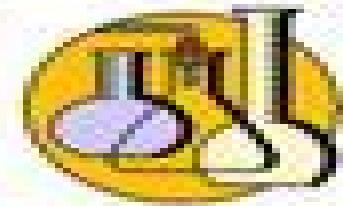


Solubility and The Mole — Reviewer Guide

Remember:

- 1 mole = 6.02×10^{23} atoms, molecules, formula units
- 1 mole = grams (of molecular mass)
- 1 mole = 22.4 liters (L) of gas



Q1. What processes need to occur when a solution that initially contained 100 g of water is diluted to 1000 g of water?

Q2. If you know the mass of a substance, you can find the number of moles by dividing it by the **molar mass of that substance**.

Q3. If you know the number of particles contained in a substance, you can find the molar mass of that substance by dividing it by...

Arranging the Problems

Q1. Summarize the steps of a grams-to-grams stoichiometry problem.

Grams A

Moles A

Moles B

Grams B

Q2. The essential stoichiometry problems involved the conversion of mass to moles and vice versa.

Moles

Moles

Q3. What reactions require a chemical equation? At the end of the reaction, chemical A will enter the products, while Chemical B has been completely used-up. Is this a Net Reacting reaction?

Chemical B

Q4. What is a mole?

A unit of measurement used to count atoms/molecules.

Q5. What is Avogadro's Number?

6.02×10^{23}

Q6. How many Avogadro's number of particles?

These numbers and values are numbers that we can use.

Mole Study Guide Key

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Moschos, William F. Griffith, Victor E.

Beshay, David D. Rahn, Debra L.

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