

Practicing the Mole -- Even Problems Answer Key

Calculate the mass in grams of each of the following:

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 26.98 \text{ g} \\ \hline \end{array} \right| = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I_2)

$$\frac{7.00 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 253.8 \text{ g} \\ \hline \end{array} \right| = 1,780 \text{ g } I_2$$

4. 2.00×10^2 moles of chlorine (Cl_2)

$$\frac{200. \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 70.90 \text{ g} \\ \hline \end{array} \right| = 14,200 \text{ g } Cl_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol}}{1 \text{ mol}} \left| \begin{array}{c} 55.85 \text{ g} \\ \hline \end{array} \right| = 514 \text{ g Fe}$$

Calculate the number of moles of atoms in each of the following:

10. 281 g of iron

$$\frac{281 \text{ g}}{55.85 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g}}{30.97 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g}}{24.31 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g}}{207.2 \text{ g}} \left| \begin{array}{c} 1 \text{ mol} \\ \hline \end{array} \right| = 4.995 \text{ mol Pb}$$

Calculate the number of atoms in each of the following:

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol}}{1 \text{ mole}} \left| \begin{array}{c} 6.022 \times 10^{23} \text{ atoms} \\ \hline \end{array} \right|$$

$$= 6.02 \times 10^{24} \text{ atoms Ti}$$

Mole Practice Problems

Michael R. Lindeburg

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