

## Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of  $\text{C}_2\text{H}_6$  and 3.44 g of  $\text{O}_2$ , determine the limiting reagent

$$14.8 \text{ g C}_2\text{H}_6 \times \frac{1 \text{ mol C}_2\text{H}_6}{26.06 \text{ g C}_2\text{H}_6} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_2\text{H}_6} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

$\text{O}_2$  is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.34 \text{ g H}_2\text{O}$$

c) determine the number of grams of  $\text{H}_2\text{O}$  produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol C}_2\text{H}_6}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g C}_2\text{H}_6}{1 \text{ mol C}_2\text{H}_6} = .0948 \text{ g C}_2\text{H}_6 \text{ used}$$

$$14.8 \text{ g C}_2\text{H}_6 - 0.0948 \text{ g C}_2\text{H}_6 = 13.9 \text{ g C}_2\text{H}_6 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{394.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g Na}_2\text{SO}_4 \times \frac{1 \text{ mol NaOH}}{204.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

# Limiting Reagent Worksheet 202 Answer Key

**SJ Ball**

**Limiting Reagent Worksheet 202 Answer Key:**

## **Limiting Reagent Worksheet 202 Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Limiting Reagent Worksheet 202 Answer Key**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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web stone age first known period of prehistoric human culture characterized by the use of stone tools the term is little used by specialists today see paleolithic period mesolithic period neolithic period stone tool industry see also bronze age iron age

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web jan 12 2018 the stone age began about 2 6 million years ago when researchers found the earliest evidence of humans using stone tools and lasted until about 3 300 b c when the bronze age began it is

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web the stone age is further subdivided by the types of stone tools in use the stone age is the first period in the three age

## **Limiting Reagent Worksheet 202 Answer Key**

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system frequently used in archaeology to divide the timeline of human technological prehistory into functional periods with the next two being the bronze age and the iron age respectively

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web early humans began using stones as simple tools about 2 million years ago humans used mainly stone tools until about 10 000 years ago however the stone age began and ended at different times in various parts of the world scientists divide the stone age into two main periods the paleolithic period and the neolithic period