

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{25.513 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{1.569 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarities of the unknown AgNO_3 solution and the unknown NaCl solution. Round your data and show your calculations.

Unknown AgNO_3

Upon the addition of solid NaCl , solid AgCl was formed.

AgCl formed (g) = 2.32019 g

$$\text{mass}(\text{AgNO}_3) = \frac{M_r(\text{AgNO}_3)}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{149.87}{143.32} \frac{\text{g}}{\text{mol}} \times 2.32019 \text{ g} = \boxed{2.3500 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{n_{\text{AgNO}_3}}{V_{\text{AgNO}_3}} = \frac{2.3500 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.16189 \frac{\text{mol}}{\text{L}}}$$

Unknown NaCl

Upon the addition of solid AgNO_3 , solid AgCl was formed.

AgCl formed (g) = 10.7907 g

$$\text{mass}(\text{NaCl}) = \frac{M_r(\text{NaCl})}{M_r(\text{AgCl})} \times \text{mass}(\text{AgCl}) = \frac{58.443}{143.32} \frac{\text{g}}{\text{mol}} \times 10.7907 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{n_{\text{NaCl}}}{V_{\text{NaCl}}} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} + \frac{1000 \text{ mL}}{1 \text{ L}} = \boxed{0.75294 \frac{\text{mol}}{\text{L}}}$$

Conclusion questions

1. If you have a company that is using this method to create AgCl , knowing that AgNO_3 is an expensive chemical, what do you think would be a better limiting reagent to use, AgNO_3 or NaCl ? Why?

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