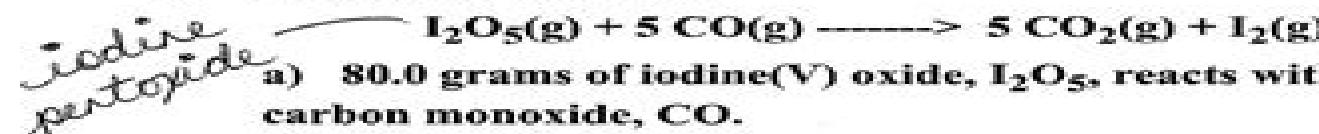


Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction



a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO .

Determine the mass of iodine I_2 , which could be produced?

b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.

i) what mass of iodine was produced?

ii) what percentage yield of iodine was produced.

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

b) How many grams of ZnS will be formed?

c) How many grams of the excess reactant will remain after the reaction is over?

3. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen?

What mass is in excess? What mass of MgO is formed?

4. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S?

5. When MoO_3 and Zn are heated together they react



What mass of ZnO is formed when 20.0 grams of MoO_3 is reacted with 10.0 grams of Zn?

6. Silver nitrate, AgNO_3 , reacts with ferric chloride, FeCl_3 , to give silver chloride, AgCl , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$. In a particular experiment, it was planned to mix a solution containing 25.0 g of AgNO_3 with another solution containing 45.0 grams of FeCl_3 .

a) Write the chemical equation for the reaction.

b) Which reactant is the limiting reactant?

c) What is the maximum number of moles of AgCl that could be obtained from this mixture?

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E Staff

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Holt Chemistry Ralph Thomas Myers, 2004 **Essentials of Chemistry** Dennis D. Staley, Antony C. Wilbraham, 1984 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and

Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application. *Basic Concepts of Chemistry, Study Guide* Leo J. Malone, 2000. The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry. Designed to help readers master the principles of general chemistry. As a prep book it promotes active involvement with the material. There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills. Updated to include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems. **Modern Chemistry**, 1990. [Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e](#) Kenneth W. Raymond, 2005-10-07. Finally readers have a shorter less intimidating introduction to general organic and biological chemistry. Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand. In this approach similarities can be exploited and concepts reinforced. The result is that readers see the strong connections that exist between these three branches of chemistry.

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