

# Limiting Reagents and Percentage Yield Worksheet

## 1. Consider the reaction

*iodine  
pentoxide*



a) 80.0 grams of iodine(V) oxide,  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ .

Determine the mass of iodine  $\text{I}_2$ , which could be produced?

b) If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.

i) what mass of iodine was produced?

ii) what percentage yield of iodine was produced.

## 2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

b) How many grams of  $\text{ZnS}$  will be formed?

c) How many grams of the excess reactant will remain after the reaction is over?

## 3. Which element is in excess when 3.00 grams of $\text{Mg}$ is ignited in 2.20 grams of pure oxygen?

What mass is in excess? What mass of  $\text{MgO}$  is formed?

## 4. How many grams of $\text{Al}_2\text{S}_3$ are formed when 5.00 grams of $\text{Al}$ is heated with 10.0 grams $\text{S}$ ?

## 5. When $\text{MoO}_3$ and $\text{Zn}$ are heated together they react



What mass of  $\text{ZnO}$  is formed when 20.0 grams of  $\text{MoO}_3$  is reacted with 10.0 grams of  $\text{Zn}$ ?

## 6. Silver nitrate, $\text{AgNO}_3$ , reacts with ferric chloride, $\text{FeCl}_3$ , to give silver chloride, $\text{AgCl}$ , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$ . In a particular experiment, it was planned to mix a solution containing 25.0 g of $\text{AgNO}_3$ with another solution containing 45.0 grams of $\text{FeCl}_3$ .

a) Write the chemical equation for the reaction.

b) Which reactant is the limiting reactant?

c) What is the maximum number of moles of  $\text{AgCl}$  that could be obtained from this mixture?

# Limiting Reagent And Percent Yield 9 3 Review Answers

**E Staff**



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### **Table of Contents Limiting Reagebt And Percent Yield 9 3 Review Answers**

1. Understanding the eBook Limiting Reagebt And Percent Yield 9 3 Review Answers
  - The Rise of Digital Reading Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagebt And Percent Yield 9 3 Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Personalized Recommendations
  - Limiting Reagebt And Percent Yield 9 3 Review Answers User Reviews and Ratings
  - Limiting Reagebt And Percent Yield 9 3 Review Answers and Bestseller Lists

5. Accessing Limiting Reagebt And Percent Yield 9 3 Review Answers Free and Paid eBooks
  - Limiting Reagebt And Percent Yield 9 3 Review Answers Public Domain eBooks
  - Limiting Reagebt And Percent Yield 9 3 Review Answers eBook Subscription Services
  - Limiting Reagebt And Percent Yield 9 3 Review Answers Budget-Friendly Options
6. Navigating Limiting Reagebt And Percent Yield 9 3 Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reagebt And Percent Yield 9 3 Review Answers Compatibility with Devices
  - Limiting Reagebt And Percent Yield 9 3 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Highlighting and Note-Taking Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Interactive Elements Limiting Reagebt And Percent Yield 9 3 Review Answers
8. Staying Engaged with Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagebt And Percent Yield 9 3 Review Answers
9. Balancing eBooks and Physical Books Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagebt And Percent Yield 9 3 Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Setting Reading Goals Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Fact-Checking eBook Content of Limiting Reagebt And Percent Yield 9 3 Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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