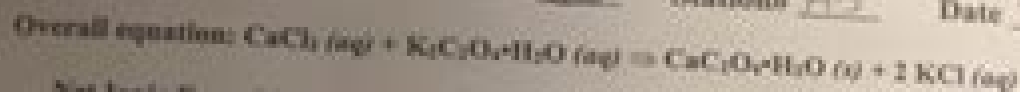


Name Mikayla D'AnnaSection 854Station A5Date 10/14/19

1. Data & Calculations:

<p>CaCl₂</p> <p>Concentration of solution <u>0.75 mol/L</u></p> <p>Volume of solution <u>15 mL</u></p> <p>Moles of CaCl₂ used <u>5.92×10^{-4} mol</u></p>	<p>K₂C₂O₄·H₂O</p> <p>Mass assigned <u>0.140 g</u></p> <p>Mass used <u>0.142 g</u></p> <p>Moles of K₂C₂O₄·H₂O used <u>0.0008 mol</u> (MW = 184.14 g/mol)</p>
<p>Crucible and filter mass</p> <p>Dry (crucible + filter) <u>22.115 g</u></p> <p>With Precipitate <u>22.221 g</u></p>	<p>CaC₂O₄·H₂O (Precipitate)</p> <p>Mass of CaC₂O₄·H₂O recovered <u>0.106 g</u></p> <p>Moles of CaC₂O₄·H₂O recovered <u>0.0007 mol</u> (MW = 146.12 g/mol)</p>

Be sure to include the necessary units for the values reported in the above table. Attach sample calculations of each type necessary to obtain the results in the above table. These should include the equations used, substituted values, appropriate units, and results.

2. Excess and Limiting Tests:

2a. Did a precipitate form (1) when K₂C₂O₄ solution was added to the filtrate? yes (2) when CaCl₂ solution was added? No

2b. Based on the results shown in 2a, what was the limiting reactant? K₂C₂O₄ What was the excess reactant? CaCl₂

Limiting Reactant Lab Report Answers

Sara McCubbins, Angela Codron



Limiting Reactant Lab Report Answers:

Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *Instructor's Guide for Introductory Chemistry in the Laboratory* James F. Hall, 1996 Working with Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

Chemistry McGraw-Hill Staff, 2001-07 *Forensics in Chemistry* Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K's body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in why chemistry concepts are relevant and how they connect All chapters include case information specific to each performance assessment and highlight the related national standards and chemistry content Chapters provide Teacher guides to help you set up Student performance assessments A suspect file to introduce the characters and new information about their relationships to the case Samples of student work that has been previously assessed and that serves as an answer key for you Grading rubrics Using Forensics in Chemistry as your guide you will gain the confidence to use inquiry based strategies and performance based assessments with a complex chemistry curriculum Your students may gain an interest in chemistry that rivals their fascination with Bones and CSI *The Student's Lab Companion* John W. Lehman, 2008 This comprehensive lab companion provides enough theory to help students understand how and why an operation works but emphasizes the practical aspects of an operation to help them perform the operation successfully in the lab For undergraduate or graduate students taking organic chemistry lab This comprehensive lab companion provides enough theory to help students understand how and why an operation works but emphasizes the practical aspects of an operation to help them perform the

operation successfully in the lab The Second Edition makes substantive revisions of many operations to clarify existing material and add new information More environmentally friendly i e green lab experiments are encouraged Ideal for professors who write their own lab experiments or would like custom labs but need a source for lab operations and safety information **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 **The Zinc and Iodine Book** Stephen

DeMeo, 2019-05-15 This book is for chemistry teachers who are thinking about reinventing their laboratory experiments that they provide to their students More than a collection of experiments it is an example of using a chemical theme to teach chemistry Instead of introducing many different chemicals per experiment as is the norm in most lab manuals this novel resource focuses on two commonly found elements Zinc and Iodine So what is so special about these elements At the heart of this resource is a colorful cyclic reaction between zinc and iodine one that produces a compound that can decompose back to its original elements This unique phenomenon demonstrates that matter not only changes but is also conserved through a chemical reaction Knowing that a compound can be the same but different than the reactants that formed it is to understand the essence of chemical change Complementing this reaction this book contains experimental activities that utilize the zinc and iodine theme to scaffold new concepts such as the properties of matter solid and gas stoichiometry equilibrium kinetics acid base chemistry and electrochemistry This teacher tested resource focuses on a set of safe substances that are appropriate for high school teachers who provide an advanced chemistry placement course and for college instructors teaching a first year chemistry laboratory sequence **Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M.

Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 *Government Reports Announcements & Index*, 1994

Scientific and Technical Aerospace Reports, 1965 **U.S. Government Research & Development Reports**, 1967

Energy Research Abstracts, 1985 **Cooperative Learning in the Chemistry Classroom** Melissa Ann Flynn, 1999

Government Reports Announcements, 1975-05-02 **Bibliography of Scientific and Industrial Reports**, 1971

Science John Michels (Journalist), 2010 U.S. Government Research & Development Reports, 1968-04 **Survival**

Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching *INIS Atomindex*, 1988

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