

$$\text{Rate}_1 = \frac{0.0001 \text{ g} - 0.00001 \text{ g}}{0.0001 \text{ mol}} = \frac{999.99 \text{ g}}{0.0001 \text{ mol}} = 9.9999 \times 10^6 \text{ g/mol}$$

4. Calculate that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

#### Part B - Unknown solutions

Using a method similar to the one used in part A, calculate and carry out a procedure for determining the molarities of the unknown Ag<sup>+</sup>(aq) solutions and the unknown Pb<sup>2+</sup>(aq) solutions. Present your data and show your calculations.

#### Unknown Ag<sup>+</sup>(aq)

Open the solution of which Pb<sup>2+</sup>, and Ag<sup>+</sup>(aq) was found.

Ag<sup>+</sup>(aq) found (g) = 0.27000000000000002

$$\text{Rate}_2 = \frac{0.0001 \text{ g} - 0.27000000000000002 \text{ g}}{0.0001 \text{ mol}} = \frac{-0.26999999999999998 \text{ g}}{0.0001 \text{ mol}} = -2.6999999999999998 \times 10^6 \text{ g/mol}$$

$$\text{Rate}_3 = \frac{0.0001 \text{ g} - 0.27000000000000002 \text{ g}}{0.0001 \text{ mol}} = \frac{-0.26999999999999998 \text{ g}}{0.0001 \text{ mol}} = -2.6999999999999998 \times 10^6 \text{ g/mol}$$

#### Unknown Pb<sup>2+</sup>(aq)

Open the solution of which Ag<sup>+</sup>(aq), and Pb<sup>2+</sup>(aq) was found.

Pb<sup>2+</sup>(aq) found (g) = 0.00010000000000000002

$$\text{Rate}_4 = \frac{0.0001 \text{ g} - 0.00010000000000000002 \text{ g}}{0.0001 \text{ mol}} = \frac{-0.00009999999999999998 \text{ g}}{0.0001 \text{ mol}} = -9.999999999999998 \times 10^3 \text{ g/mol}$$

$$\text{Rate}_5 = \frac{0.0001 \text{ g} - 0.00010000000000000002 \text{ g}}{0.0001 \text{ mol}} = \frac{-0.00009999999999999998 \text{ g}}{0.0001 \text{ mol}} = -9.999999999999998 \times 10^3 \text{ g/mol}$$

#### Conclusion

5. If you have a company that is using this method to remove Ag<sup>+</sup>(aq), knowing that Ag<sup>+</sup>(aq) is an essential (when do you think) would the a better cleaning method to use, Ag<sup>+</sup>(aq) or Pb<sup>2+</sup>(aq)?

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**Marcel A. Müller**

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