

**Limiting Reagent & Percent Yield Practice Worksheet**

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

**Only 13 g can be formed**

c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.482 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu}(\text{NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu}(\text{NO}_3)_2}{1 \text{ mol Cu}(\text{NO}_3)_2} = 20.9 \text{ g Cu}(\text{NO}_3)_2$$

e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.482 \text{ g CuCl}_2} \times 2 \text{ mol NaNO}_3 \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb}(\text{NO}_3)_2 \times \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{331.214 \text{ g Pb}(\text{NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb}(\text{NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

**8.51 g NaNO<sub>3</sub> can be made**

# Limiting Reagent And Percentage Yield Review Key

**James E. Brady,Fred Senese**

## **Limiting Reagent And Percentage Yield Review Key:**

*Introduction to General, Organic, and Biochemistry* Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject *Introduction to General Organic and Biochemistry* 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

**Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

**Addison-Wesley Chemistry**, 1990 **MCAT 528 Advanced Prep**

**2023-2024 Kaplan Test Prep**, 2022-11-01 Kaplan s MCAT 528 Advanced Prep 2023 2024 features thorough subject review more questions than any competitor and the highest yield questions available all authored by the experts behind Kaplan s score raising MCAT prep course Prepping for the MCAT is a true challenge Kaplan can be your partner along the way offering guidance on where to focus your efforts how to organize your review and targeted focus on the most tested concepts This edition features commentary and instruction from Kaplan s MCAT experts and has been updated to match the AAMC s guidelines precisely no more worrying if your MCAT review is comprehensive The Most Practice More than 500 questions in the book and online and access to even more online more practice than any other advanced MCAT book on the market The Best Practice Comprehensive subject review is written by top rated award winning Kaplan instructors All material is vetted by editors with advanced science degrees and by a medical doctor Online resources including a full length practice test help you master the computer based format you ll see on Test Day Expert Guidance Star Ratings throughout the book indicate how important each topic will be to your score on the real exam informed by Kaplan s decades of MCAT experience and facts straight from the testmaker We know the test The Kaplan MCAT team has spent years studying every MCAT related document available Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test

*5 Steps to a 5 AP Chemistry, 2010-2011 Edition* John Moore, Richard H. Langley, 2010-01-01 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You ll sharpen your

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ways to think of skills as tools that can be used to solve complex problems. Several aids are included to help focus and inspire student interest: frequent reference to common chemicals in commercial products; numerous photographs of reactions in chapter practice exercises following worked examples.

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