

Limiting Reagent and Percentage Yield

1 Dichlorine monoxide is produced by passing chlorine gas over heated mercury (II) oxide according to the following unbalanced reaction:



Balance the reaction. Suppose the quantity of the reactants is sufficient to produce 0.86 g of Cl_2O , but only 0.71 g is obtained. What is the percentage yield?

2 Consider the following reaction:



The theoretical yield of beryllium chloride was 10.7 g. Suppose the reaction actually yields 4.5 g. What is the percentage yield?

3 Nitrogen and hydrogen at high temperatures are converted to ammonia using the following reaction:



When 400 g of H_2 are added to an excess amount of N_2 , 104 g of NH_3 are formed. Calculate the percentage yield.

Limiting Reagent And Percent Yield Workbook

George Facer

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The MCAT Chemistry Book Ajikumar Aryangat, 2021-01-15 Comprehensive Rigorous Prep for MCAT Chemistry The MCAT Chemistry Book presents a comprehensive review of general chemistry and organic chemistry to prepare for the Medical College Admission Test Part I presents general chemistry concepts and Part II presents organic chemistry concepts The review sections are written in a user friendly manner to simplify and reduce the student's burden when deciphering difficult concepts At the end of each chapter practice questions are included to test the understanding of the key concepts Answers and explanations for the practice questions are provided after the review sections Illustrations and tables are included wherever necessary to focus and clarify key ideas and concepts

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