

## Limiting Reagents and Percentage Yield Worksheet

### 1. Consider the reaction

*iodine  
pentoxide*



a) 80.0 grams of iodine(V) oxide,  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ .

Determine the mass of iodine  $\text{I}_2$ , which could be produced?

b) If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.

i) what mass of iodine was produced?

ii) what percentage yield of iodine was produced.

### 2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

b) How many grams of  $\text{ZnS}$  will be formed?

c) How many grams of the excess reactant will remain after the reaction is over?

### 3. Which element is in excess when 3.00 grams of $\text{Mg}$ is ignited in 2.20 grams of pure oxygen?

What mass is in excess? What mass of  $\text{MgO}$  is formed?

### 4. How many grams of $\text{Al}_2\text{S}_3$ are formed when 5.00 grams of $\text{Al}$ is heated with 10.0 grams $\text{S}$ ?

### 5. When $\text{MoO}_3$ and $\text{Zn}$ are heated together they react



What mass of  $\text{ZnO}$  is formed when 20.0 grams of  $\text{MoO}_3$  is reacted with 10.0 grams of  $\text{Zn}$ ?

### 6. Silver nitrate, $\text{AgNO}_3$ , reacts with ferric chloride, $\text{FeCl}_3$ , to give silver chloride, $\text{AgCl}$ , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$ . In a particular experiment, it was planned to mix a solution containing 25.0 g of $\text{AgNO}_3$ with another solution containing 45.0 grams of $\text{FeCl}_3$ .

a) Write the chemical equation for the reaction.

b) Which reactant is the limiting reactant?

c) What is the maximum number of moles of  $\text{AgCl}$  that could be obtained from this mixture?

# Limiting Reagent And Percent Yield 9 3 Section Review

**Angelica M. Stacy**



### **Limiting Reagent And Percent Yield 9 3 Section Review:**

**Chemistry: Core Concepts, 3rd Edition** Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package

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**Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

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*Living by Chemistry (2018 Update)* Angelica M. Stacy,2019-03-07 Designed to help all students to learn chemistry Living by Chemistry is a full year high school curriculum that incorporates science practices with a guided inquiry approach Students of all levels will gain a deep understanding of chemistry with this program With Living by Chemistry students learn chemistry in the same way that chemists work by asking questions collecting evidence and thinking like scientists Living by Chemistry is the product of a decade of research and development in high school classrooms focusing on optimizing student understanding of chemical principles Author Angelica Stacy assisted in the development of the NGSS standards and served on the AP Chemistry redesign committee She designed Living by Chemistry as an introduction for students who will take AP Chemistry or additional college classes The curriculum was developed with the belief that science is best learned through first hand experience and discussion with peers Guided inquiry allows students to actively participate in and become adept at scientific processes and communication These skills are vital to a students further success in science as well as beneficial to other pursuits Formal definitions and formulas are frequently introduced after students have explored

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