

# ICE Practice Problems

**#1 – Relatively easy, no ICE table required because eq'm concentrations are given**

For the reaction  $\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}(g) + 3\text{H}_2(g)$  @  $1500^\circ\text{C}$  an equilibrium mixture of these gases was found to have the following concentrations:  $[\text{CO}] = 0.300\text{M}$ ,  $[\text{H}_2] = 0.800\text{M}$  and  $[\text{CH}_4] = 0.400\text{M}$ .  $K_c @ 1500^\circ\text{C} = 5.67$ . Determine the equilibrium concentration of  $\text{H}_2\text{O}$  in this mixture.

**#2 – Requires an ICE table because you do not know the equilibrium concentrations- no product is yet formed**

For the reaction  $\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$  calculate the equilibrium concentrations of all species if 1.000 mol of each reactant is mixed in a 1.000L flask.  $K_c = 5.10$  at the temperature of this reaction.



[Initial]	I	1.000	1.000	0	0
[Change in]	C				
[Equilibrium]	E				

**#3 – Requires an ICE table because you do not know the equilibrium concentrations. The initial concentrations must be calculated- no product is yet formed**

For the reaction  $\text{H}_2(g) + \text{F}_2(g) \rightleftharpoons 2\text{HF}(g)$  calculate the equilibrium concentrations of all species if 3.000 mol of each reactant was added 1.500L flask.  $K_c$  at the temperature of the reaction is  $1.15 \times 10^2$ .



[Initial]	I	2.00	2.00	0	
[Change in]	C				
[Equilibrium]	E				

**#4 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants are given.**

0.200mol of  $\text{H}_2$ , 0.200mol of  $\text{I}_2$ , and 0.200mol of  $\text{HI}$  were placed in a 1.00 L flask and allowed to come to equilibrium. The  $K_c$  value of the reaction at this temperature is 49.5. Determine the equilibrium concentrations of all species.



[Initial]	I				
[Change in]	C				
[Equilibrium]	E				

**#5 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants must be calculated and no product is yet formed**

For the reaction  $\text{H}_2(g) + \text{F}_2(g) \rightarrow 2\text{HF}(g)$  calculate the equilibrium concentrations of each species if 3.000 mol of  $\text{H}_2$  and 6.000mol of  $\text{F}_2$  are mixed in a 3.000L flask.  $K_c$  at this temperature is  $1.15 \times 10^2$ .



[Initial]	I				
[Change in]	C				
[Equilibrium]	E				

# Ice Practice Problems Answer Key

**Prentice-Hall Staff**

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