

2**MATTER AND CHANGE****SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)**

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The mass of an object is a measure of the amount of matter the object contains.
2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a substance.
4. Is the following sentence true or false? All samples of a substance have different physical properties. false
5. A physical property is a quality or condition of a substance that can be observed or measured without changing the substance's composition.
6. Circle the letter of the term that is NOT a physical property.
 - a. hardness
 - b. color
 - c. boiling point
 - d. melting
7. Look at Table 2.1 on page 40. What is the melting point of bromine? -7°C
8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.
 - a. sulfur
 - b. chlorine
 - c. gold
 - d. copper

Manual Chemistry Matter And Change Answer Key

Katrin Zwirglmaier

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Chemistry Laurel Dingrado,McGraw-Hill Staff,Glencoe/McGraw-Hill,2001-08-01 **Chemistry** McGraw-Hill Staff,2001-07 **Quanta, Matter and Change: A Molecular Approach to Physical Change** Peter Atkins,Julio de Paula,Ron Friedman,2008-11-15 Beginning with quantum mechanics introducing statistical mechanics and progressing through to thermodynamics this new text for the two semester physical chemistry course features a wealth of new applications and insights as well as new Mathematical Background inter chapters to help students review key quantitative concepts This is a splendid book True to the authors philosophy as outlined in the preface it approaches physical chemistry by first developing the quantum theory of molecular electronic structure then by statistical arguments moves into thermodynamics and thence to kinetics Peter Taylor Review in Chemistry World Royal Society of Chemistry July 31 2009

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