

Electron Configuration Practice

Chemistry

How to write an electron configuration:

- A. Determine the total number of electrons to be represented.
- B. Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- C. The sum of the superscripts should equal the total number of electrons. Example:
 $_{12}Mg \ 1s^2 \ 2s^2 \ 2p^6 \ 3s^2$

I. Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the pr per number of electrons for the element) are filled.

1. Na: $1s^2 \ 2s^2 \ 2p^6 \ 3s^1$
2. Pb: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10} \ 6p^2$
3. Sr: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2$
4. U: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10} \ 6p^6 \ 7s^2 \ 5f^3 \ 6d^1$
5. N: $1s^2 \ 2s^2 \ 2p^3$
6. Ag: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^1 \ 4d^{10}$
7. Ti: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^2$
8. Ce: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^1 \ 5d^1$
9. Cl: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^5$
10. Hg: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10}$

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Kathy A. Zahler

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