

Answer key-H.w(3)-Electrochemical cells

Multiple choice questions:

1- What is electrolysis?

g-is the decomposition of a compound into its elements by an electric current

b-making a compound using a current

2-What is an electrolyte?

a-The liquid part of the electrochemical cell

b-The metal part of the electrochemical cell

c-The part that measures the voltage in an electrochemical cell

3-Where does oxidation occur?

a-Cathode

6 - Anode

e-Electrode

4- What way do electrons flow

α -From species being oxidised to species being reduced

b—From species being reduced to species being oxidised

c-They don't flow

5-Which statement correctly describes the 2 electrodes?

a-The anode is negative and the cathode is positive

b-The anode and cathode are both positive

c-The anode is positive and the cathode is negative

6- What is the name given to the solution that is being electrolysed?

a-Salt solution

b-Electric solution

c-Electrolyte

7-Explain why the electrolyte has to be a liquid.

So the ions can move towards the electrodes

b-So the electrons can move

c-The current can flow through it.

B- What happens at the positive electrode?

a-Positive non-metal ions are attracted

b-Positive metal ions are attracted

c-Negative anions (non-metal) ions are attracted

9- When NaCl solution is electrolysed what has forms at the anode?

2-convexity

b-Chlorine

c-sodium

10-What does the Cathode (-) do to ions?

a-Give electrons to the Positive ions to turn them back into atoms

b-Take electrons from the positive ions?

c-Turn ions back into atoms by removing electrons

11- What is the equation to show what happens to Cl ions at the anode.

$$a \cdot d' \rightarrow d + e'$$
$$b-2Cl^- \rightarrow Cl_2 + 2e^-$$
$$\text{Fe}(\text{OH})_3 + 3\text{e}^- \rightarrow 3\text{OH}^-$$

Electrochemical Cells Answer Key

Charles H. Corwin



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