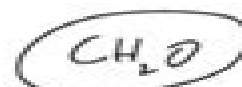


# Empirical and Molecular Formulas Worksheet 1

- The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.

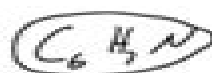
$$\frac{39.9\%}{12.01} \times \frac{1}{3.32} = 1 \quad \frac{6.7\%}{1.008} \times \frac{1}{2.32} = 1$$



- The molar mass for question #1 was determined by experiment to be 60.0 g/mol. What is the molecular formula?

$$\frac{60.0}{30.03} = 2 \quad \text{C}_2\text{H}_4\text{O}_2$$

- Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{N}_2$  as products. If the combustion of 9.71 g of aniline yields 6.63 g  $\text{H}_2\text{O}$  and 1.46 g  $\text{N}_2$ , what is its empirical formula?



$$\frac{6.63\text{g}}{18.02} \times \frac{1}{1.12} = 0.736 \text{ mol H} \quad \frac{1.46\text{g}}{28.02} \times \frac{1}{1.12} = 0.104 \text{ mol N}$$

- The molar mass of aniline is 93 g/mol. What is its molecular formula?

$$\frac{93}{92.13} = 1 \quad \text{C}_6\text{H}_7\text{N}$$

- Calculate the mass percent of carbon, nitrogen and oxygen in acetamide,  $\text{C}_2\text{H}_5\text{NO}$ .

$$\frac{2(12.01)}{59.07} = 40.66\% \text{ C} \quad \frac{5(1.008)}{59.07} = 8.532\% \text{ H} \quad \frac{1(14.01)}{59.07} = 23.72\% \text{ N} \quad \frac{1(16.00)}{59.07} = 27.07\% \text{ O}$$

- A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?

$$\frac{11.39\text{g P}}{30.97} \times \frac{1}{1} = 0.3678 \text{ mol P} \quad \frac{50.51 - 11.39}{35.45} \times \frac{1}{1} = 1.103 \text{ mol Cl}$$



# Empirical And Molecular Formula Answer Key

**Ralph Thomas Myers**



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