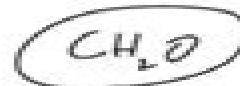


Empirical and Molecular Formulas Worksheet 1

1. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.

$$\frac{39.9\% C}{12.01} = 3.32 \sim 1 \quad \frac{6.7\% H}{1.008} = 6.67 \sim 1$$

$$\frac{53.4\% O}{16.00} = 3.34 \sim 1 \quad \frac{3.34 \sim 1}{3.32} = 1$$



a. The molar mass for question #1 was determined by experiment to be 60.0 g/mol. What is the molecular formula?

$$\frac{60.0}{30.0} = 2$$

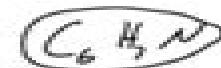


2. Aniline, a starting material for urethane plastic foams, consists of C, H, and N.

Combustion of such compounds yields  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{N}_2$  as products. If the combustion of 9.71 g of aniline yields 6.63 g  $\text{H}_2\text{O}$  and 1.46 g  $\text{N}_2$ , what is its empirical formula?

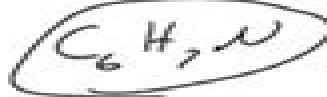
$$\frac{6.63\text{g H}_2\text{O}}{18.02\text{g}} \frac{1-16}{1-16.0} = \frac{.732 \sim 1\text{H}}{1.00} = 7 \quad \frac{1.46\text{g N}_2}{28.02\text{g}} \frac{1-14}{1-14.0} = \frac{1.71 \sim 1\text{N}}{1.00} = 1.71 \sim 1.6 = 2.20 \text{g N}_2 \sim 2.20 \text{g N}_2 = 7.56\text{g C}$$

$$\frac{9.71\text{g C}}{12.01\text{g}} \frac{1-12}{1-12.0} = \frac{.804 \sim 1\text{C}}{1.00} = 1 \quad \frac{7.56\text{g C}}{12.01\text{g}} \frac{1-1}{1-12.01} = \frac{.625 \sim 1\text{C}}{1.00} = 6$$



a. The molar mass of aniline is 93 g/mol. What is its molecular formula?

$$\frac{93}{93.13} = 1$$



3. Calculate the mass percent of carbon, nitrogen and oxygen in acetamide,  $\text{C}_2\text{H}_5\text{NO}$ .

$$\frac{2(12.01)}{59.07} \quad \frac{5(1.008)}{59.07} \quad \frac{1(14.01)}{59.07} \quad \frac{1(16.00)}{59.07}$$

$$40.66\% \text{C} \quad 8.532\% \text{H} \quad 22.72\% \text{N} \quad 27.07\% \text{O}$$

4. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?

$$\frac{11.39\text{g P}}{30.97} = \frac{.3678 \sim 1}{.3678} = 1$$

$$50.51 - 11.39 = 39.12 \text{g Cl} \quad \frac{1-35.45}{1-35.45} = \frac{1.03 \sim 1}{.3678} = 3$$



# Empirical And Molecular Formula Answer Key

**Ralph Thomas Myers**

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