

1. Consider two acids HCO_2H (formic acid, $\text{Pka} = 3.8$) and pivalic acid $[(\text{CH}_3)_3\text{CCO}_2\text{H}]$, $\text{Pka} = 5.0$ (4)

a. Which acid has the larger K_a ?

pivalic acid

formic acid

$\text{Pka} = 5.0$

$\text{pka} = 3.8$

$$\begin{aligned} K_a &= 10^{-\text{pka}} \\ &= 10^{-5.0} \\ &= 10^{-5} \end{aligned}$$

(0.5)

$$\begin{aligned} K_a &= 10^{-3.8} \\ &= 1.5 \times 10^{-4} \end{aligned}$$

Formic acid has higher K_a (0.5)

b. Which acid is strong acid?

Formic acid is strong (less Pka) (1)

c. Which acid forms the stronger conjugate base?

Pivalic acid will form a strong conjugate base (1)

d. When each acid is dissolved in water, for which acid does the equilibrium lie further to the right? Illustrate your reasoning with a reaction scheme.



Grademathematics Tutorial Memorandum

Annelies Wilder-Smith



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Table of Contents Grademathematics Tutorial Memorandum

1. Understanding the eBook Grademathematics Tutorial Memorandum
 - The Rise of Digital Reading Grademathematics Tutorial Memorandum
 - Advantages of eBooks Over Traditional Books
2. Identifying Grademathematics Tutorial Memorandum
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grademathematics Tutorial Memorandum
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grademathematics Tutorial Memorandum
 - Personalized Recommendations
 - Grademathematics Tutorial Memorandum User Reviews and Ratings

- Grademathematics Tutorial Memorandum and Bestseller Lists
- 5. Accessing Grademathematics Tutorial Memorandum Free and Paid eBooks
 - Grademathematics Tutorial Memorandum Public Domain eBooks
 - Grademathematics Tutorial Memorandum eBook Subscription Services
 - Grademathematics Tutorial Memorandum Budget-Friendly Options
- 6. Navigating Grademathematics Tutorial Memorandum eBook Formats
 - ePub, PDF, MOBI, and More
 - Grademathematics Tutorial Memorandum Compatibility with Devices
 - Grademathematics Tutorial Memorandum Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grademathematics Tutorial Memorandum
 - Highlighting and Note-Taking Grademathematics Tutorial Memorandum
 - Interactive Elements Grademathematics Tutorial Memorandum
- 8. Staying Engaged with Grademathematics Tutorial Memorandum
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grademathematics Tutorial Memorandum
- 9. Balancing eBooks and Physical Books Grademathematics Tutorial Memorandum
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grademathematics Tutorial Memorandum
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grademathematics Tutorial Memorandum
 - Setting Reading Goals Grademathematics Tutorial Memorandum
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grademathematics Tutorial Memorandum
 - Fact-Checking eBook Content of Grademathematics Tutorial Memorandum
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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