

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the molar mass of the excess reactant.

4 Followed moles of product \times mass of product

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 400 grams of oxygen, determine the limiting reactant.

6 How many grams of water produced?

7 Followed moles of product \times mass of product



8 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
 If you start with 10 moles of magnesium and 2 moles of sodium, determine the limiting reactant.

9 Calculate the grams of magnesium chloride produced.

10 How many grams of the excess reactant is produced in a left over?

11 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
 If you have 100 grams of magnesium and 400 grams of sodium, determine the limiting reactant.

12 How much sodium chloride is produced?

13 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 400 grams of oxygen, determine the limiting reactant.

14 Calculate the grams of carbon dioxide produced.
 $CO_2 = 44.01 \text{ g/mol}$
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180.16 \text{ g}} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times 44.01 \text{ g/mol} = 146.4 \text{ g } CO_2$

15 How many grams of water produced?
 $H_2O = 18.02 \text{ g/mol}$
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180.16 \text{ g}} \times \frac{6 \text{ mol } H_2O}{1 \text{ mol } C_6H_{12}O_6} \times 18.02 \text{ g/mol} = 108.1 \text{ g } H_2O$

16 How many moles of carbon dioxide is produced?
 $CO_2 = 44.01 \text{ g/mol}$
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180.16 \text{ g}} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} = 3.33 \text{ mol } CO_2$

17 How many grams of the excess reactant is produced in a left over?
 $O_2 = 32.00 \text{ g/mol}$
 $400 \text{ g } O_2 - 146.4 \text{ g } CO_2 = 253.6 \text{ g } O_2$

18 How many grams of the excess reactant is produced in a left over?
 $H_2O = 18.02 \text{ g/mol}$
 $400 \text{ g } H_2O - 108.1 \text{ g } H_2O = 291.9 \text{ g } H_2O$

19 How many moles of carbon dioxide is produced?
 $CO_2 = 44.01 \text{ g/mol}$
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180.16 \text{ g}} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} = 3.33 \text{ mol } CO_2$

20 How many grams of the excess reactant is produced in a left over?
 $O_2 = 32.00 \text{ g/mol}$
 $400 \text{ g } O_2 - 146.4 \text{ g } CO_2 = 253.6 \text{ g } O_2$

ANSWER KEY INCLUDED

Limiting Reactant Problems Answer Key

Stephen DeMeo



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