

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K?

$$P = 2.3 \text{ atm}$$

$$n = \frac{PV}{RT}$$

$$n = \frac{2.3 \cdot 120}{0.0821 \cdot 340}$$

$$9.887 \text{ moles}$$

$$V = 120 \text{ L}$$

$$T = 340 \text{ K}$$

2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?

$$P = ?$$

$$V = 50 \text{ L}$$

$$T = 200^\circ \text{C} = 473 \text{ K}$$

$$n = 45 \text{ moles}$$

$$P = \frac{45 \cdot 0.0821 \cdot 473}{50}$$

$$34.95 \text{ atm}$$

3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?

$$V = 1.0 \text{ L}$$

4) $n = ?$
How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?

5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3 moles of oxygen gas at a pressure of 1 atmosphere, what is the temperature of the balloon?

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CL Gary

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