

12.3 Limiting Reagent and Percent Yield > Limiting and Excess Reagents

A balanced chemical equation is a chemist's recipe.

Chemical Equations					
	$\text{N}_2(g)$	+	$3\text{H}_2(g)$	$\longrightarrow$	$2\text{NH}_3(g)$
"Microscopic recipe"	1 molecule N_2	+	3 molecules H_2	$\longrightarrow$	2 molecules NH_3
"Macroscopic recipe"	1 mol N_2	+	3 mol H_2		2 mol NH_3

- What would happen if two molecules (moles) of N_2 reacted with three molecules (moles) of H_2 ?

Experimental Conditions		
	Reactants	Products
Before reaction	 2 molecules N_2	 3 molecules H_2
		0 molecules NH_3

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Morris Hein, Susan Arena



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