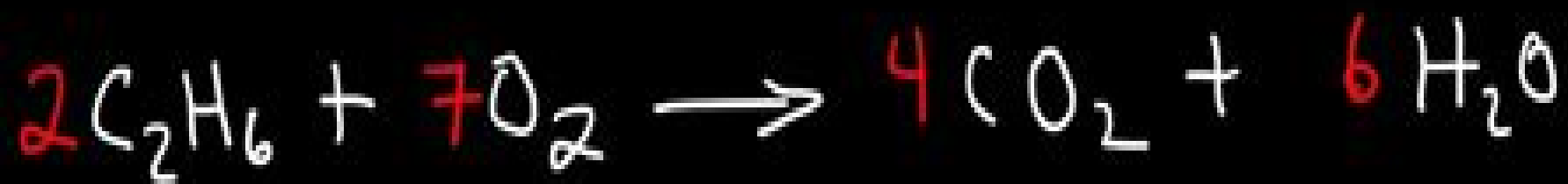


Limiting Reactant



$$\frac{84\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32\text{g O}_2} \times \frac{6\text{mol H}_2\text{O}}{7\text{mol O}_2} \times \frac{18.016\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}}$$

$$= 40.54\text{g H}_2\text{O}$$

Limiting Reactants C1y Vm

Lifelique



Limiting Reactants C1y Vm:

Limiting Reactant and Percent Lifelique, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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