

Limiting Reagent and Percentage Yield

- 1 Dichlorine monoxide is produced by passing chlorine gas over heated mercury (II) oxide according to the following unbalanced reaction:



Balance the reaction. Suppose the quantity of the reactants is sufficient to produce 0.86 g of Cl_2O , but only 0.71 g is obtained. What is the percentage yield?

- 2 Consider the following reaction:



The theoretical yield of beryllium chloride was 10.7 g. Suppose the reaction actually yields 4.5 g. What is the percentage yield?

- 3 Nitrogen and hydrogen at high temperatures are converted to ammonia using the following reaction:



When 400 g of H_2 are added to an excess amount of N_2 , 104 g of NH_3 are formed. Calculate the percentage yield.

Limiting Reagent And Percent Yield Answers

RJ Alexander



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

Access to Chemistry Avril Higton, Mike Clemmet, Elaine Golding, Alan V Jones, 2007-10-31 This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening

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